

Chapter 17 Thermochemistry Study Guide

17.1 – 17.2 Thermochemical Equations

General Chemistry

1. Make the following conversions:
 - a. 444 cal to joules = 1.86×10^3 J
 - b. 1.8 kJ to joules = 1.8×10^3 J
 - c. 0.45 kJ to calories = 1.1×10^2 cal
2. Classify each of these processes as endothermic or exothermic:
 - a. condensing steam – **exo**
 - b. burning alcohol – **exo**
 - c. evaporating alcohol – **endo**
 - d. baking a potato – **endo**
3. The specific heat capacity for silver is $0.24 \text{ J/g}^\circ\text{C}$. Calculate the energy required to raise the temperature of 150.0 g Ag from 273 K to 298 K. Calculate the molar heat capacity of silver.

$$\text{Energy: } q = (150.0)(.24)(298-273) = 9.0 \times 10^2 \text{ J}$$

$$\text{Molar heat capacity: } (0.24 \text{ J/g}^\circ\text{C})(197.87 \text{ g/mol Ag}) = 26 \text{ J/mol}$$

4. It takes 585 J of energy to raise the temperature of 125.6 g Hg from 20.0°C to 53.5°C . Calculate the specific heat capacity and the molar heat capacity of Hg.

$$\text{Specific heat capacity: } C = q/m\Delta t = 585/(125.6)(53.5-20) = 0.139 \text{ J/g}^\circ\text{C}$$

$$\text{Molar heat capacity: } (0.139 \text{ J/g}^\circ\text{C})(200.59 \text{ g/mol Hg}) = 27.9 \text{ J/mol}$$

5. A 46.2-g sample of copper is heated to 95.4°C and then placed in a calorimeter containing 75.0 g water at 19.6°C . The equilibrium temperature in the calorimeter is 21.8°C . Calculate the specific heat capacity of copper, assuming that all the heat lost by the copper is gained by the water.

$$C = 0.203 \text{ J/g}^\circ\text{C}$$

6. A 15.0-g sample of nickel metal is heated to 100.0°C and dropped into 55.0 g of water, initially at 23.0°C . Assuming that all the heat lost by the nickel is absorbed by the water; calculate the final temperature of the nickel and the water. The specific heat of nickel is $0.444 \text{ J/g}^\circ\text{C}$.

$$\begin{aligned} -15.0(444)(T_f - 100.0) &= 55.0(4.184)(T_f - 23.0) \\ -6.66T_f + 666 &= 230.12T_f - 5292.76 \\ 5956.76 &= 236.78 T_f \\ &= T_f \\ &= 25.2^\circ\text{C} \end{aligned}$$

7. Chloe was running bath water and realized it was too hot. If she has 20.0 L of water in the tub at 95°C and then adds 15.8 L of water at 75°C , what will the final temperature of the water be?

$$\begin{aligned} -20(T_f - 95) &= 15.8(T_f - 75) \\ -20T_f + 1900 &= 15.8 T_f - 1185 \\ 3085 &= 35.8 T_f \end{aligned}$$

Chapter 17 Thermochemistry Interpreting Graphics

Answers

Huangqi Zhang



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Table of Contents Chapter 17 Thermochemistry Interpreting Graphics Answers

1. Understanding the eBook Chapter 17 Thermochemistry Interpreting Graphics Answers
 - The Rise of Digital Reading Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 17 Thermochemistry Interpreting Graphics Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Personalized Recommendations
 - Chapter 17 Thermochemistry Interpreting Graphics Answers User Reviews and Ratings
 - Chapter 17 Thermochemistry Interpreting Graphics Answers and Bestseller Lists
5. Accessing Chapter 17 Thermochemistry Interpreting Graphics Answers Free and Paid eBooks
 - Chapter 17 Thermochemistry Interpreting Graphics Answers Public Domain eBooks
 - Chapter 17 Thermochemistry Interpreting Graphics Answers eBook Subscription Services
 - Chapter 17 Thermochemistry Interpreting Graphics Answers Budget-Friendly Options
6. Navigating Chapter 17 Thermochemistry Interpreting Graphics Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chapter 17 Thermochemistry Interpreting Graphics Answers Compatibility with Devices
 - Chapter 17 Thermochemistry Interpreting Graphics Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Highlighting and Note-Taking Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Interactive Elements Chapter 17 Thermochemistry Interpreting Graphics Answers
 8. Staying Engaged with Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 17 Thermochemistry Interpreting Graphics Answers
 9. Balancing eBooks and Physical Books Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 17 Thermochemistry Interpreting Graphics Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Setting Reading Goals Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Fact-Checking eBook Content of Chapter 17 Thermochemistry Interpreting Graphics Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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