

## Chapter 12

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_\_.
- |   |  |
|---|--|
| a. add the coefficients of the reagents | c. convert given quantities to volumes |
| b. convert given quantities to moles    | d. convert given quantities to masses  |
- \_\_\_\_\_ 2. In the reaction  $2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g)$ , what is the ratio of moles of oxygen used to moles of  $\text{CO}_2$  produced?
- |        |        |
|--------|--------|
| a. 1:1 | c. 1:2 |
| b. 2:1 | d. 2:2 |
- \_\_\_\_\_ 3. How many moles of glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , can be "burned" biologically when 10.0 mol of oxygen is available?
- $$\text{C}_6\text{H}_{12}\text{O}_6(l) + 6\text{O}_2(g) \rightarrow 6\text{CO}_2(g) + 6\text{H}_2\text{O}(l)$$
- |              |             |
|--------------|-------------|
| a. 0.500 mol | c. 33.3 mol |
| b. 1.67 mol  | d. 60.0 mol |
- \_\_\_\_\_ 4. At STP, how many liters of oxygen are required to react completely with 3.0 liters of hydrogen to form water?
- $$2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(g)$$
- |          |          |
|----------|----------|
| a. 1.5 L | c. 2.0 L |
| b. 3.0 L | d. 2.4 L |
- \_\_\_\_\_ 5. The equation below shows the decomposition of lead nitrate. How many grams of oxygen are produced when 11.5 g  $\text{NO}_2$  is formed?
- $$2\text{Pb}(\text{NO}_3)_2(s) \rightarrow 2\text{PbO}(s) + 4\text{NO}_2(g) + \text{O}_2(g)$$
- |           |           |
|-----------|-----------|
| a. 1.00 g | c. 2.88 g |
| b. 2.00 g | d. 32.0 g |
- \_\_\_\_\_ 6. Iron(III) oxide is formed when iron combines with oxygen in the air. How many grams of  $\text{Fe}_2\text{O}_3$  are formed when 16.7 g of Fe reacts completely with oxygen?
- $$4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$$
- |           |           |
|-----------|-----------|
| a. 32.0 g | c. 47.8 g |
| b. 23.9 g | d. 85.6 g |
- \_\_\_\_\_ 7. When glucose is consumed, it reacts with oxygen in the body to produce carbon dioxide, water, and energy. How many grams of carbon dioxide would be produced if 45 g of  $\text{C}_6\text{H}_{12}\text{O}_6$  completely reacted with oxygen?
- |          |         |
|----------|---------|
| a. 1.5 g | c. 11 g |
| b. 1.8 g | d. 66 g |
- \_\_\_\_\_ 8. Aluminum reacts with sulfuric acid to produce aluminum sulfate and hydrogen gas. How many grams of aluminum sulfate would be formed if 250 g  $\text{H}_2\text{SO}_4$  completely reacted with aluminum?
- $$2\text{Al}(s) + 3\text{H}_2\text{SO}_4(aq) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$$
- |           |          |
|-----------|----------|
| a. 0.85 g | c. 450 g |
| b. 290 g  | d. 870 g |

# Chapter 12 Test M Lingerfelts Blog

**Huangqi Zhang**



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