

Chapter 12

SECTION 12.1

12.1.1 Boyle's Law: Pressure-Volume Relationship for an Ideal Gas

1. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph.

Graph 1: A graph showing pressure (y-axis) versus volume (x-axis) for a fixed mass of gas at constant temperature. The curve is a hyperbola, indicating an inverse relationship between pressure and volume.

Graph 2: A graph showing pressure (y-axis) versus volume (x-axis) for a fixed mass of gas at constant temperature. The curve is a straight line, indicating a direct relationship between pressure and volume.

Graph 3: A graph showing pressure (y-axis) versus volume (x-axis) for a fixed mass of gas at constant temperature. The curve is a straight line, indicating a direct relationship between pressure and volume.

Graph 4: A graph showing pressure (y-axis) versus volume (x-axis) for a fixed mass of gas at constant temperature. The curve is a hyperbola, indicating an inverse relationship between pressure and volume.

2. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph.

3. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph.

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7. Boyle's Law states that pressure of a fixed mass of gas is inversely proportional to its volume. This is the relationship shown in the following graph.

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SA Dillow



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