

## Chemistry Student Edition - Basic Answer Key

### Chapter 11: Chemical Reactions

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#### Chemical Equations

##### Questions

1. Early Chinese alchemists discovered an early form of gunpowder. What was the composition of this substance?
2. What later developments were made to the gunpowder recipe that improved its pyrotechnic properties?
3. Make an argument for why the burning of a candle is consistent with the law of conservation of matter/mass.
4. Think of an experiment that you could conduct to demonstrate that mass is conserved for a given chemical change.
5. Balance the following chemical equations:
  - a.  $C + O_2 \rightarrow CO$
  - b.  $CO + O_2 \rightarrow CO_2$
  - c.  $H_2 + Br_2 \rightarrow HBr$
  - d.  $K + H_2O \rightarrow KOH + H_2$
  - e.  $O_3 \rightarrow O_2$
  - f.  $N_2 + H_2 \rightarrow NH_3$
  - g.  $Zn + AgCl \rightarrow ZnCl_2 + Ag$
  - h.  $Cl_2 + NaI \rightarrow NaCl + I_2$
  - i.  $P_4O_{10} + H_2O \rightarrow H_3PO_4$
  - j.  $Be_2C + H_2O \rightarrow Be(OH)_2 + CH_4$
  - k.  $S + HNO_3 \rightarrow H_2SO_4 + NO_2 + H_2O$
  - l.  $NH_3 + CuO \rightarrow Cu + N_2 + H_2O$
  - m.  $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

##### Answers

1. Early gunpowder was comprised of saltpeter (impure potassium nitrate), sulfur, and charcoal (carbon).
2.
  - a. Purification of potassium nitrate
  - b. Using potassium chlorate instead of potassium nitrate
  - c. Further refinement of the recipe to find the proportions of reactants that yielded the most energetic reaction
3. The burning of a candle obeys the law of conservation of mass because, even though the mass of the candle will decrease as it burns, the mass is not lost. It is transferred into carbon dioxide and water that enter the surrounding air.
4. The reaction between iron and oxygen to produce rust could be studied as an example of conservation of mass. Steel wool could be placed into a covered beaker containing water. The reaction would proceed over time and rust would begin to form. The beaker could be weighed before and after the change. The mass would be the same, even though a new

# Chapter 11 Chemical Reactions Vocabulary Review

## Answer Key

**Robert C. Smoot, Smoot, Richard G.  
Smith, Jack Price**



## Chapter 11 Chemical Reactions Vocabulary Review Answer Key:

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Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions including clear explanations and step by step problem solving. The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans rules, collision number and mean free path, abnormal molecular masses and van t Hoff factor and applications of radioactivity. **Chemistry: Core Concepts, 3rd Edition** Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09  
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**Investigating Chemistry** Matthew Johll, 2008-12-22 In its new second edition *Investigating Chemistry A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package *Scott Foresman Science*, 2008 **McGraw-Hill's SAT Subject Test: Chemistry, 2ed** Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test We've put all of our proven expertise into McGraw Hill's SAT Subject Test Chemistry to make sure you're fully prepared for this difficult exam With this book you'll get essential skill building techniques and strategies created by a leading high school chemistry teacher You'll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill's SAT Subject Test Chemistry we'll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score

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**Glencoe Science**, 2002. Physical Science George A. Williams, Richard Barnes, Ray Doerhoff, Max Bolen, 1979. A physical science text stressing an awareness of the environment with related laboratory activities to lead the student into discovering basic laws and concepts of physics and chemistry. **Understanding Chemistry** Fred M. Dewey, 1994

**Focus on physical science** Charles H. Heimler, 1989. *Chemistry Textbook for College and University USA* Ibrahim Sikder, 2023-06-04. **Chemistry** John W. Moore, Conrad L. Stanitski, Peter C. Jurs, 2002. CHEMISTRY: THE MOLECULAR SCIENCE is intended to help students develop a broad overview of chemistry and chemical reactions, an understanding of the most important concepts and models that chemists and those in chemistry-related fields use, an appreciation of the many ways chemistry impacts our daily lives, the ability to apply the facts, concepts, and models of chemistry appropriately to new situations in chemistry, other sciences, and engineering, and to other disciplines.

Introductory Chemistry Steven S. Zumdahl, 1993. **Matter and Energy**, 2005

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