

CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY

1) Which of the following are weak electrolytes? HCl, $\text{HC}_2\text{H}_3\text{O}_2$, NH_3 , KCl

A) HCl, KCl

B) HCl, $\text{HC}_2\text{H}_3\text{O}_2$, NH_3 , KCl

C) $\text{HC}_2\text{H}_3\text{O}_2$, KCl

D) $\text{HC}_2\text{H}_3\text{O}_2$, NH_3

E) HCl, $\text{HC}_2\text{H}_3\text{O}_2$, KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and HNO_3 (aq)?

A) K^+ and H^+

C) K^+ and NO_3^-

E) OH^- only

B) H^+ and OH^-

D) H^+ and NO_3^-

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of CaCO_3 when aqueous solutions of Na_2CO_3 and CaCl_2 are mixed is _____.

A) $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Na}_2\text{CO}_3(\text{aq})$

E)

B) $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq})$

$\text{Na}_2\text{CO}_3(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{CaCO}_3(\text{s})$

C) $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{aq})$

D) $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$

Answer: D Diff: 3

4) Which of the following are strong acids? HI, HNO_3 , HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI, HNO_3 , HBr

B) HI, HNO_3 , HF, HBr

D) HNO_3 , HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces _____.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

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McGraw-Hill Staff



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General Chemistry Ralph H. Petrucci, Robert K. Wismer, 1989

The Chemical News and Journal of Industrial Science, 1882

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Table of Contents Chemistry Matter Change Chapter 8 Answer Key

1. Understanding the eBook Chemistry Matter Change Chapter 8 Answer Key
 - The Rise of Digital Reading Chemistry Matter Change Chapter 8 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter Change Chapter 8 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter Change Chapter 8 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter Change Chapter 8 Answer Key

- Personalized Recommendations
 - Chemistry Matter Change Chapter 8 Answer Key User Reviews and Ratings
 - Chemistry Matter Change Chapter 8 Answer Key and Bestseller Lists
5. Accessing Chemistry Matter Change Chapter 8 Answer Key Free and Paid eBooks
- Chemistry Matter Change Chapter 8 Answer Key Public Domain eBooks
 - Chemistry Matter Change Chapter 8 Answer Key eBook Subscription Services
 - Chemistry Matter Change Chapter 8 Answer Key Budget-Friendly Options
6. Navigating Chemistry Matter Change Chapter 8 Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - Chemistry Matter Change Chapter 8 Answer Key Compatibility with Devices
 - Chemistry Matter Change Chapter 8 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Matter Change Chapter 8 Answer Key
 - Highlighting and Note-Taking Chemistry Matter Change Chapter 8 Answer Key
 - Interactive Elements Chemistry Matter Change Chapter 8 Answer Key
8. Staying Engaged with Chemistry Matter Change Chapter 8 Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter Change Chapter 8 Answer Key
9. Balancing eBooks and Physical Books Chemistry Matter Change Chapter 8 Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter Change Chapter 8 Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Matter Change Chapter 8 Answer Key
- Setting Reading Goals Chemistry Matter Change Chapter 8 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Matter Change Chapter 8 Answer Key

- Fact-Checking eBook Content of Chemistry Matter Change Chapter 8 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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