

Stoichiometry Worksheet 2
Chemistry

Name _____

Date _____ Hour _____

1. In the reaction : $2\text{Al}_2\text{O}_{3(s)} \rightarrow 4\text{Al}_{(s)} + 3\text{O}_{2(g)}$, the mole ratios are:

Al_2O_3 to Al _____ Al_2O_3 to O_2 _____ Al to O_2 _____ O_2 to Al _____

2. In the reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, the mole ratio of oxygen to water is: _____

3. In the reaction: $\text{Ca} + \text{Cl}_2 \rightarrow \text{CaCl}_2$, the mole ratio of chlorine to calcium is: _____

4. In the reaction: $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$, the mole ratio of zinc to sulfuric acid is:

5. The equation for the Haber process is represented by:



To convert 9.0 moles of hydrogen to ammonia would require how many moles of nitrogen?

6. In the equation: $2\text{KClO}_{3(g)} \rightarrow 2\text{KCl}_{(s)} + 3\text{O}_{2(g)}$, how many moles of oxygen are produced when 3.00 moles of KClO_3 decompose completely?

7. For the reaction: $\text{C}_{(s)} + 2\text{H}_{2(g)} \rightarrow \text{CH}_{4(g)}$, how many moles of H_2 are required to produce 10.0 moles of CH_4 ?

Chapter 9 Stoichiometry Section 2 Worksheet

Robert R. Redfield



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