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chapter - Chemical Kinetics - 01.

Admit
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- i) Instantaneous Rate of Reaction.
- (a) It occurs within a short span of time.
 - (b) It cannot be calculated for multistep ~~elementary~~ reaction.
 - (c) It can be calculated for elementary reaction.

- Average Rate of Reaction.
- (a) It occurs during a long interval of time.
 - (b) It can be calculated for multistep reaction.
 - (c) It cannot be calculated for elementary reaction.

ii) If the unit of rate constant is s^{-1} , the reaction is of first order.

iii) No, the molecularity can never be zero or a fractional number as it shows the number of reactants taking part in a reaction can never be zero.

iv)



Rate of rxn (M) = $k[R]^n$ → order of rxn. Reaction involves two reactants.

? $x = k[R]^n$ — (1)

$8x = k[2R]^n$ — (2)

Dividing eq (2) by (1);

$$\frac{8x}{x} = \frac{k[2R]^n}{k[R]^n}$$

or, $8 = 2^n \cdot \frac{R^n}{R^n}$

Overall order

= (a+b)

= (1+2) = 3

or, $8 = 2^n$

or, $2^3 = 2^n$

$\therefore n = 3$

The order of reaction is 3.

(1) $y = k[A]^a[B]^b$

(2) $2y = k[2A]^a[2B]^b$

(3) $8y = k[2A]^a[2B]^b$

(2) ÷ (1) →

$$\frac{2y}{y} = \frac{k[2A]^a[2B]^b}{k[A]^a[B]^b}$$

$\sim 2^1 = 2^a \cdot 2^b$ or $a=1$

(3) ÷ (2) →

$2^2 = 4 = 2^2 \cdot 2^b$ or $b=2$

(v) The rate constant for the first order reaction is $60 s^{-1}$.

let M be the molar mass.
a be the initial concentration.

$k = 60/s$

final concentration = $\frac{M}{16} \times a$

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Christian Drosten



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