

## Introduction 11

# An Electron's Address

Name: \_\_\_\_\_  
 Date: \_\_\_\_\_  
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### Introduction: Energy Levels and Sublevels

In addition to the energy level model, Bohr proposed the following model of energy levels. The energy level of the atom can be thought of as the distance between the nucleus and the energy level. The distance between the nucleus and the energy level is the energy level.

Each energy level contains one or more sublevels. Draw the following table:

Table 1

Energy Level	Sublevels in each energy level
1	s
2	s, p
3	s, p, d
4	s, p, d, f

Draw the table in words. Draw as a structural model of the  $1^{\text{st}}$  energy level. Draw the model of the  $2^{\text{nd}}$  energy level. Draw the model of the  $3^{\text{rd}}$  energy level.

### Critical Thinking Questions

- How many sublevels are in the  $1^{\text{st}}$  energy level?  
 The sublevel in the  $1^{\text{st}}$  energy level is the s sublevel. It is the only sublevel in the  $1^{\text{st}}$  energy level.
- How many sublevels are in the  $2^{\text{nd}}$  energy level?  
 The sublevels in the  $2^{\text{nd}}$  energy level are the s and p sublevels.
- How many sublevels are in the  $3^{\text{rd}}$  energy level?  
 The sublevels in the  $3^{\text{rd}}$  energy level are the s, p, and d sublevels.
- How many sublevels would you expect to have for the  $4^{\text{th}}$  energy level?  
 The number of sublevels in the  $4^{\text{th}}$  energy level would be 4 sublevels. There would be s, p, d, and f sublevels.
- How many sublevels would you expect to have for the  $5^{\text{th}}$  energy level?  
 The number of sublevels in the  $5^{\text{th}}$  energy level would be 5 sublevels. There would be s, p, d, and f sublevels.

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