

Nomenclature Review

Name: \_\_\_\_\_

1. Give the charge and name for each of the following *oxyanions*.

	Charge	Name
a. NO <sub>3</sub>	-1	nitrate
b. CO <sub>3</sub>	-2	carbonate
c. C <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	-1	acetate
d. ClO <sub>2</sub>	-1	chlorite
e. PO <sub>3</sub>	-3	phosphite
f. ClO <sub>4</sub>	-1	perchlorate
g. SO <sub>4</sub>	-2	sulfate
h. NH <sub>4</sub>	+1	ammonium
i. OH	-1	hydroxide
j. CN	-1	cyanoide
k. BrO <sub>3</sub>	-1	bromate
l. C <sub>2</sub> O <sub>4</sub>	-2	oxalate
m. SO <sub>3</sub>	-2	sulfate
n. MnO <sub>4</sub>	-1	permanganate

2. If A and B represent compounds of an element bonded to a polyatomic ion, give an example of a valid formula for each combination below.

a. AB	b. AB <sub>2</sub>	c. A <sub>2</sub> B	d. AB <sub>3</sub>	e. A <sub>2</sub> B <sub>3</sub>
Al(NO <sub>3</sub> ) <sub>3</sub>	Ca(NO <sub>2</sub> ) <sub>2</sub>	(NH <sub>4</sub> ) <sub>2</sub> S	Al(ClO <sub>4</sub> ) <sub>3</sub>	Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>

3. What is the oxidation number of:

a. Calcium ion?	+2		
b. Sodium ion?	+1		
c. Sulfur ion?	-2		
d. Phosphorus ion?	-3		
e. Carbon ion within Carbon dioxide molecule?	+4	$x + 2(-2) = 0$ $x = 4+$	
f. Phosphorus within the Phosphate polyatomic ion?	+5		$x + 4(-2) = -2$ $x = +6$
g. Chromium within the Chromate polyatomic ion?	+6		
h. Carbon within the Oxalate polyatomic ion?	+3		$x + 4(-2) = -2$ $2x + 4(-2) = -2$ $x = 3+$
i. Zinc ion?	+2		
j. Cadmium ion?	+2		
k. Sulfur within the Hydrogen Sulfite ion (HSO <sub>3</sub> ) <sup>-</sup>	+4	$1 + x - 6 = -1$ $x = 4+$	

4. Give the most common oxidation number/numbers for each of the following groups or families of elements.

a. Alkali Metals	+1
b. Alkaline Earth	+2
c. Noble Gases	0
d. Halogens	-1
e. Group 13	+3
f. Group 15	-3

5. Give the formula for the following hydrocarbons/alcohols:

a. methane	C <sub>1</sub> H <sub>4</sub>
b. butene	C <sub>4</sub> H <sub>8</sub>
c. ethanol	C <sub>2</sub> H <sub>5</sub> OH
d. propyne	C <sub>3</sub> H <sub>4</sub>

# Chemistry Unit 4 Test Review Electron Configuration

**RC Schank**

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## **Table of Contents Chemistry Unit 4 Test Review Electron Configuration**

1. Understanding the eBook Chemistry Unit 4 Test Review Electron Configuration
  - The Rise of Digital Reading Chemistry Unit 4 Test Review Electron Configuration
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 4 Test Review Electron Configuration
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Unit 4 Test Review Electron Configuration
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 4 Test Review Electron Configuration

- Personalized Recommendations
  - Chemistry Unit 4 Test Review Electron Configuration User Reviews and Ratings
  - Chemistry Unit 4 Test Review Electron Configuration and Bestseller Lists
5. Accessing Chemistry Unit 4 Test Review Electron Configuration Free and Paid eBooks
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    - Chemistry Unit 4 Test Review Electron Configuration Budget-Friendly Options
  6. Navigating Chemistry Unit 4 Test Review Electron Configuration eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Unit 4 Test Review Electron Configuration Compatibility with Devices
    - Chemistry Unit 4 Test Review Electron Configuration Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Unit 4 Test Review Electron Configuration
    - Highlighting and Note-Taking Chemistry Unit 4 Test Review Electron Configuration
    - Interactive Elements Chemistry Unit 4 Test Review Electron Configuration
  8. Staying Engaged with Chemistry Unit 4 Test Review Electron Configuration
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Unit 4 Test Review Electron Configuration
  9. Balancing eBooks and Physical Books Chemistry Unit 4 Test Review Electron Configuration
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Unit 4 Test Review Electron Configuration
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Unit 4 Test Review Electron Configuration
    - Setting Reading Goals Chemistry Unit 4 Test Review Electron Configuration
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Unit 4 Test Review Electron Configuration

- Fact-Checking eBook Content of Chemistry Unit 4 Test Review Electron Configuration
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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