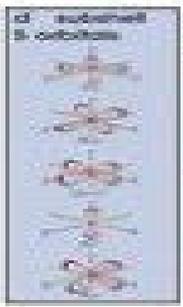


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PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box indicating its position: **1.008** - atomic weight, **H** - atomic symbol, **1** - atomic number.



ATOMIC STRUCTURE

Atomic Number, Z , is of protons in the nucleus. For neutral atom, $Z = +$ of electrons. For ions of charge, $+Q$, $Z = +Q$ (cations) or $Z = -Q$ (anions). Atomic mass, A , is the sum of protons and neutrons in the nucleus. Isotopes atoms with the same Z , different A . The atomic mass has a number of isotopes. Use atomic mass, a number of isotopes, as practical work use the atomic mass. Number of Atomic Weight is practical average of atomic masses for a given element.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number (n) - determines the size and energy of the orbital. $n = 1, 2, 3, \dots$
Angular momentum quantum number (l) - determines the shape of the orbital. $l = 0, 1, 2, \dots, n-1$
Magnetic quantum number (m_l) - determines the orientation of the orbital. $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$
Spin quantum number (m_s) - determines the spin of the electron. $m_s = +\frac{1}{2}, -\frac{1}{2}$

NUCLEAR CHEMISTRY

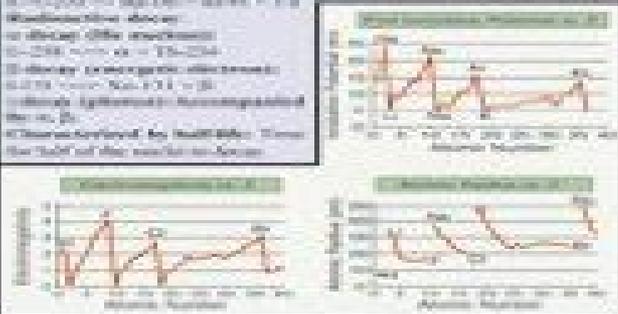
Nuclear reactions alter the number, particles present or chemical behavior of atoms and have energy. Energy (E) is related to mass (m) by the Einstein equation: $E = mc^2$.
 Number of protons in nucleus, Z , number is symbolized by ${}^A_Z X$, or ${}^A_Z X$ (e.g. Deuterium, ${}^2_1\text{D}$ or ${}^2_1\text{H}$).
 Types of processes:
 - **Radioactivity:** Two electrons made from particle collisions.
 - **Alpha decay:** ${}^4_2\text{He}$ or ${}^4_2\text{He}^{2+}$
 - **Beta decay:** ${}^0_{-1}\text{e}$ or ${}^0_{-1}\text{e}^{-}$
 - **Gamma decay:** ${}^0_0\gamma$ or ${}^0_0\gamma$
 - **Neutron capture:** ${}^1_0\text{n}$ or ${}^1_0\text{n}^0$
 - **Proton capture:** ${}^1_1\text{p}$ or ${}^1_1\text{p}^+$
 - **Electron capture:** ${}^0_{-1}\text{e}$ or ${}^0_{-1}\text{e}^{-}$
 - **Positron emission:** ${}^0_{+1}\text{e}$ or ${}^0_{+1}\text{e}^{+}$
 - **Alpha decay:** ${}^4_2\text{He}$ or ${}^4_2\text{He}^{2+}$
 - **Beta decay:** ${}^0_{-1}\text{e}$ or ${}^0_{-1}\text{e}^{-}$
 - **Gamma decay:** ${}^0_0\gamma$ or ${}^0_0\gamma$
 - **Neutron capture:** ${}^1_0\text{n}$ or ${}^1_0\text{n}^0$
 - **Proton capture:** ${}^1_1\text{p}$ or ${}^1_1\text{p}^+$
 - **Electron capture:** ${}^0_{-1}\text{e}$ or ${}^0_{-1}\text{e}^{-}$
 - **Positron emission:** ${}^0_{+1}\text{e}$ or ${}^0_{+1}\text{e}^{+}$

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe all many-electron atoms and ions. A many electron atom has many different energy levels and orbital properties. The main point is that orbitals have different energies. The filling of levels is guided by the Aufbau Principle.

n	l	m _l	m _s	Orbitals
1	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1s
2	0	0	$+\frac{1}{2}, -\frac{1}{2}$	2s
2	1	0, ± 1	$+\frac{1}{2}, -\frac{1}{2}$	2p
3	0	0	$+\frac{1}{2}, -\frac{1}{2}$	3s
3	1	0, ± 1	$+\frac{1}{2}, -\frac{1}{2}$	3p
3	2	0, $\pm 1, \pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	3d
4	0	0	$+\frac{1}{2}, -\frac{1}{2}$	4s
4	1	0, ± 1	$+\frac{1}{2}, -\frac{1}{2}$	4p
4	2	0, $\pm 1, \pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	4d
4	3	0, $\pm 1, \pm 2, \pm 3$	$+\frac{1}{2}, -\frac{1}{2}$	4f

Atomic configuration: Shows orbital occupancy. Pauli Exclusion Principle: Each orbital has a unique set of quantum numbers. An orbital may hold up to two electrons, one with spin up, $+\frac{1}{2}$, and one with spin down, $-\frac{1}{2}$.
Slater's Rules: For estimating the energy of an orbital. The energy of an orbital is estimated by the sum of the energy of the orbital and the energy of the other orbitals. The energy of an orbital is estimated by the sum of the energy of the orbital and the energy of the other orbitals. The energy of an orbital is estimated by the sum of the energy of the orbital and the energy of the other orbitals.



Block	Group	Element	Symbol	Atomic Number		
s-block	1	H	H	1		
		Li	Li	3		
	2	Na	Na	11		
		K	K	19		
	p-block	13	B	B	5	
			Al	Al	13	
		14	C	C	6	
			Si	Si	14	
		15	N	N	7	
			P	P	15	
		16	O	O	8	
			S	S	16	
		17	F	F	9	
			Cl	Cl	17	
		18	Ne	Ne	10	
			Ar	Ar	18	
		d-block	3	Sc	Sc	21
				Ti	Ti	22
4	V		V	23		
	Cr		Cr	24		
5	Mn		Mn	25		
	Fe		Fe	26		
6	Co		Co	27		
	Ni		Ni	28		
7	Cu		Cu	29		
	Zn		Zn	30		
f-block	7	La	La	57		
		Ce	Ce	58		
	8	Pr	Pr	59		
		Nd	Nd	60		
	9	Pm	Pm	61		
		Sm	Sm	62		
	10	Eu	Eu	63		
		Gd	Gd	64		
	11	Tb	Tb	65		
		Dy	Dy	66		
12	Ho	Ho	67			
	Er	Er	68			
13	Tm	Tm	69			
	Yb	Yb	70			
14	Lu	Lu	71			
	Hf	Hf	72			

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