

## 13

## ELECTRONS IN ATOMS

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- Extension topics

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- Core honors content
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SECTION 13.1 MODELS OF THE ATOM (pages 361–366) 

*This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.*

## ► The Evolution of Atomic Models (pages 361–363)

1. What are the chemical properties of atoms, ions, and molecules related to?

The properties are related to the arrangement of electrons.

2. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

3. The energy level of an electron is the region around the nucleus where the electron is likely to be moving.
4. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false
5. Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
  - maintain an electron in its present energy level
  - move an electron from its present energy level to the next higher one
6. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

# Chemistry Chapter 13 Electrons In Atoms

**CL Gary**



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### **Table of Contents Chemistry Chapter 13 Electrons In Atoms**

1. Understanding the eBook Chemistry Chapter 13 Electrons In Atoms
  - The Rise of Digital Reading Chemistry Chapter 13 Electrons In Atoms
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 13 Electrons In Atoms
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 13 Electrons In Atoms
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 13 Electrons In Atoms
  - Personalized Recommendations
  - Chemistry Chapter 13 Electrons In Atoms User Reviews and Ratings

- Chemistry Chapter 13 Electrons In Atoms and Bestseller Lists
- 5. Accessing Chemistry Chapter 13 Electrons In Atoms Free and Paid eBooks
  - Chemistry Chapter 13 Electrons In Atoms Public Domain eBooks
  - Chemistry Chapter 13 Electrons In Atoms eBook Subscription Services
  - Chemistry Chapter 13 Electrons In Atoms Budget-Friendly Options
- 6. Navigating Chemistry Chapter 13 Electrons In Atoms eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 13 Electrons In Atoms Compatibility with Devices
  - Chemistry Chapter 13 Electrons In Atoms Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 13 Electrons In Atoms
  - Highlighting and Note-Taking Chemistry Chapter 13 Electrons In Atoms
  - Interactive Elements Chemistry Chapter 13 Electrons In Atoms
- 8. Staying Engaged with Chemistry Chapter 13 Electrons In Atoms
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 13 Electrons In Atoms
- 9. Balancing eBooks and Physical Books Chemistry Chapter 13 Electrons In Atoms
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 13 Electrons In Atoms
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 13 Electrons In Atoms
  - Setting Reading Goals Chemistry Chapter 13 Electrons In Atoms
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 13 Electrons In Atoms
  - Fact-Checking eBook Content of Chemistry Chapter 13 Electrons In Atoms
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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