

$n_1 \rightarrow$  Lyman (1) (UV Region)  
 $n_2 \rightarrow$  Balmer (2) (visible)  
 $n_3 \rightarrow$  3 Paschen  
 $n_4 \rightarrow$  4 Brackett  
 $n_5 \rightarrow$  5 p-fund.

Time period =  $\frac{2\pi r}{v}$   $\propto \frac{n^3}{Z^2}$

Frequency =  $\frac{1}{T}$   $\propto \frac{Z^2}{n^3}$

No. of spectral lines =  $\frac{(n_2 - n_1)(n_2 - n_1 + 1)}{2}$

$\Delta E \propto \frac{1}{n^2}$

No. of spectral lines if  $n=1 = \frac{n(n-1)}{2}$

No. of spectral lines.  
 Lyman =  $n-1$ , Balmer =  $n-2$   
 Paschen =  $n-3$ , Brackett =  $n-4$   
 p-fund =  $n-5$

Energy of  $e^- \rightarrow n+1$

Photoelectric effect.  
 $E_{photon} = h\nu_0 + KE$   
 $h\nu_p = h\nu_0 + KE$

Stopping potential  $\Rightarrow E_{photon} - W_0 = qV$

$2\pi r = n\lambda$

No. of waves made in  $n^{th}$  orbit =  $n$

de-Broglie's Eq.

$\lambda = \frac{h}{p} \Rightarrow \lambda = \frac{h}{\sqrt{2m \cdot KE}} \quad \lambda = \frac{h}{\sqrt{2m \cdot qV}} \quad \lambda_e = \sqrt{\frac{150}{V}} \text{ \AA} = \frac{12.3}{\sqrt{V}} \text{ \AA}$

$\Delta x \Delta p \geq \frac{h}{4\pi} \quad \Delta n \cdot m \Delta v \geq \frac{h}{4\pi} \quad \Delta \theta \Delta \phi \geq \frac{h}{4\pi}$ 
  
 $\theta \rightarrow$  angular disp.  $\phi \rightarrow$  momentum.

No. of subshells in  $n^{th}$  shell or no. of orbitals  $l$  values from = 0 to  $(n-1)$

# Chemistry Note For Iit Jee

**R Sandford**



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## **Table of Contents Chemistry Note For Iit Jee**

1. Understanding the eBook Chemistry Note For Iit Jee
  - The Rise of Digital Reading Chemistry Note For Iit Jee
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Note For Iit Jee
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Note For Iit Jee
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Note For Iit Jee

- Personalized Recommendations
  - Chemistry Note For Iit Jee User Reviews and Ratings
  - Chemistry Note For Iit Jee and Bestseller Lists
5. Accessing Chemistry Note For Iit Jee Free and Paid eBooks
    - Chemistry Note For Iit Jee Public Domain eBooks
    - Chemistry Note For Iit Jee eBook Subscription Services
    - Chemistry Note For Iit Jee Budget-Friendly Options
  6. Navigating Chemistry Note For Iit Jee eBook Formats
    - ePub, PDF, MOBI, and More
    - Chemistry Note For Iit Jee Compatibility with Devices
    - Chemistry Note For Iit Jee Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Note For Iit Jee
    - Highlighting and Note-Taking Chemistry Note For Iit Jee
    - Interactive Elements Chemistry Note For Iit Jee
  8. Staying Engaged with Chemistry Note For Iit Jee
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Note For Iit Jee
  9. Balancing eBooks and Physical Books Chemistry Note For Iit Jee
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Note For Iit Jee
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Note For Iit Jee
    - Setting Reading Goals Chemistry Note For Iit Jee
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Note For Iit Jee

- Fact-Checking eBook Content of Chemistry Note For Iit Jee
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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