

Section 12.1

OBJECTIVE 1

Identify suitable Lewis acid-base adducts from the species provided:

1. BF_3
 - (a) NH_3
 - (b) H_2O
 - (c) CH_3OH
 - (d) C_2H_6
 - (e) C_2H_4
 - (f) C_2H_2
 - (g) $\text{C}_2\text{H}_6\text{O}$
 - (h) $\text{C}_2\text{H}_6\text{S}$
 - (i) $\text{C}_2\text{H}_6\text{N}_2$
 - (j) $\text{C}_2\text{H}_6\text{N}_2\text{O}$
 - (k) $\text{C}_2\text{H}_6\text{N}_2\text{O}_2$
 - (l) $\text{C}_2\text{H}_6\text{N}_2\text{O}_3$
 - (m) $\text{C}_2\text{H}_6\text{N}_2\text{O}_4$
 - (n) $\text{C}_2\text{H}_6\text{N}_2\text{O}_5$
 - (o) $\text{C}_2\text{H}_6\text{N}_2\text{O}_6$
 - (p) $\text{C}_2\text{H}_6\text{N}_2\text{O}_7$
 - (q) $\text{C}_2\text{H}_6\text{N}_2\text{O}_8$
 - (r) $\text{C}_2\text{H}_6\text{N}_2\text{O}_9$
 - (s) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{10}$
 - (t) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{11}$
 - (u) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{12}$
 - (v) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{13}$
 - (w) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{14}$
 - (x) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{15}$
 - (y) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{16}$
 - (z) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{17}$
 - (aa) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{18}$
 - (ab) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{19}$
 - (ac) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{20}$
 - (ad) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{21}$
 - (ae) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{22}$
 - (af) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{23}$
 - (ag) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{24}$
 - (ah) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{25}$
 - (ai) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{26}$
 - (aj) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{27}$
 - (ak) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{28}$
 - (al) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{29}$
 - (am) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{30}$
 - (an) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{31}$
 - (ao) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{32}$
 - (ap) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{33}$
 - (aq) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{34}$
 - (ar) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{35}$
 - (as) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{36}$
 - (at) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{37}$
 - (au) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{38}$
 - (av) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{39}$
 - (aw) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{40}$
 - (ax) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{41}$
 - (ay) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{42}$
 - (az) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{43}$
 - (ba) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{44}$
 - (bb) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{45}$
 - (bc) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{46}$
 - (bd) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{47}$
 - (be) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{48}$
 - (bf) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{49}$
 - (bg) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{50}$
 - (bh) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{51}$
 - (bi) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{52}$
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 - (bx) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{67}$
 - (by) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{68}$
 - (bz) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{69}$
 - (ca) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{70}$
 - (cb) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{71}$
 - (cc) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{72}$
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 - (cg) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{76}$
 - (ch) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{77}$
 - (ci) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{78}$
 - (cj) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{79}$
 - (ck) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{80}$
 - (cl) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{81}$
 - (cm) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{82}$
 - (cn) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{83}$
 - (co) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{84}$
 - (cp) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{85}$
 - (cq) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{86}$
 - (cr) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{87}$
 - (cs) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{88}$
 - (ct) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{89}$
 - (cu) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{90}$
 - (cv) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{91}$
 - (cw) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{92}$
 - (cx) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{93}$
 - (cy) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{94}$
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 - (dd) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{99}$
 - (de) $\text{C}_2\text{H}_6\text{N}_2\text{O}_{100}$

2. Write the Lewis structure for the following chemical equation:



a. Balance the equation:



b. Write the Lewis structure of H_3O^+ or OH^- .

c. Is addition of H_2O to H_3O^+ an exothermic reaction? Is it endothermic?

d. Write the Lewis structure of H_3O^+ or OH^- as a Lewis acid or Lewis base.

e. Write the Lewis structure of H_3O^+ or OH^- as a Lewis acid or Lewis base.

Chapter 9 Stoichiometry Section 1 Answers

R Pring



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