

Chemistry Chapter 8 Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ___ 1. What type of bond would you expect in a compound of oxygen and sodium?
 a. ionic
 b. metallic
 c. covalent
 d. polar covalent
- ___ 2. The symbols δ^- and δ^+ indicate which of the following?
 a. electronegativity
 b. electrostatic force
 c. charge
 d. partial charge
- ___ 3. Which of the following represents a nonpolar covalent molecule?
 a. NaCl
 b. Br₂
 c. HCN
 d. Au
- ___ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called _____.
 a. electron affinity
 b. polarity
 c. electronegativity
 d. ionization energy
- ___ 5. The polarity of a bond depends on the difference in the atoms' _____.
 a. electron affinities
 b. electronegativities
 c. atomic radii
 d. ionization energies
- ___ 6. What type of bond would you expect to form between nitrogen and potassium?
 a. ionic
 b. covalent
 c. polar covalent
 d. metallic
- ___ * High electronegativity differences are found in _____ bonds.
 a. metallic
 b. ionic metal - nonmetal
 c. covalent
 d. Pauling
- ___ * To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?
 a. lose
 b. gain
 c. share
 d. This cannot be determined from the given information.
- ___ 9. Which of the following does not exist as a diatomic molecule?
 a. argon
 b. hydrogen
 c. chlorine
 d. oxygen
 H N O F Cl Br I
 Gas: H₂ N₂ O₂ F₂ Cl₂ Br₂ I₂
- ___ 10. Which of these is not a covalent molecule?
 a. F₂
 b. CO
 c. LiF
 d. NH₃
- ___ 11. A dash in a Lewis structure represents _____.
 a. one electron
 b. two electrons
 c. four electrons
 d. unpaired electrons
- ___ 12. The central atom in a Lewis structure is usually the atom _____.

Chapter 5 Chemistry Test Answers

Thomas A. Evangelist



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Table of Contents Chapter 5 Chemistry Test Answers

1. Understanding the eBook Chapter 5 Chemistry Test Answers
 - The Rise of Digital Reading Chapter 5 Chemistry Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 5 Chemistry Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 5 Chemistry Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 5 Chemistry Test Answers
 - Personalized Recommendations
 - Chapter 5 Chemistry Test Answers User Reviews and Ratings
 - Chapter 5 Chemistry Test Answers and Bestseller Lists
5. Accessing Chapter 5 Chemistry Test Answers Free and Paid eBooks
 - Chapter 5 Chemistry Test Answers Public Domain eBooks
 - Chapter 5 Chemistry Test Answers eBook Subscription Services
 - Chapter 5 Chemistry Test Answers Budget-Friendly Options
6. Navigating Chapter 5 Chemistry Test Answers eBook Formats

- ePub, PDF, MOBI, and More
- Chapter 5 Chemistry Test Answers Compatibility with Devices
- Chapter 5 Chemistry Test Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 5 Chemistry Test Answers
 - Highlighting and Note-Taking Chapter 5 Chemistry Test Answers
 - Interactive Elements Chapter 5 Chemistry Test Answers
- 8. Staying Engaged with Chapter 5 Chemistry Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 5 Chemistry Test Answers
- 9. Balancing eBooks and Physical Books Chapter 5 Chemistry Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 5 Chemistry Test Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 5 Chemistry Test Answers
 - Setting Reading Goals Chapter 5 Chemistry Test Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 5 Chemistry Test Answers
 - Fact-Checking eBook Content of Chapter 5 Chemistry Test Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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