

Learning Module 2 Forward Commitment and Contingent Claim Features and Instruments

- ✓ **LO1:** Define forward contracts, futures contracts, swaps, options (calls and puts), and credit derivatives and compare their basic characteristics.
- ✓ **LO2:** Determine the value at expiration and profit from a long or a short position in a call or put option.
- ✓ **LO3:** Contrast forward commitments with contingent claims.

Forwards, Futures, and Swaps

- ✓ **LO1:** Define forward contracts, futures contracts, swaps, options (calls and puts), and credit derivatives and compare their basic characteristics.

Exhibit 1 Derivative contract comparison

	Forward commitments	Contingent claims
Contract	Exchange of obligations	Consequence of rights
Counterparty	Both obligated to perform	Buyer has the right, but no obligation, to exercise; seller's obligation is contingent on exercise
Cost at initiation	Usually no cost to either party	Buyer pays premium to seller
Examples	Futures, forwards, most swaps	Options, credit default swaps, total return swaps

Derivative contracts are designated as either forward commitment contracts or contingent claims.

A forward commitment contract is an exchange of obligations between the counterparties. Examples of forward commitment contracts include:

- Forward or futures contracts: Counterparties consent to exchange an asset for a specified price at a specified time in the future. A bond/futures contract is a forward commitment.
- Fixed-for-floating interest rate swaps: Counterparties consent to exchange a stream of fixed interest payments from the fixed-rate payer for a stream of floating interest payments from the floating-rate receiver.

Contingent claims are derivatives that give holders the right to trade the underlying asset at a specified price at a future date. Contract holders do not have to exercise that right, but if they do, the original contract seller is obligated to take the other side of the trade that results from the exercise. For contingent claims, sellers must hold their contracts contingent on the contract holders exercising their rights.

Forwards

Exhibit 2 Forward commitment derivatives

	Futures	Forward	Swaps
Trading venue	Exchange-traded	Over-the-counter	Over-the-counter
Legal counterparty	Exchange clearinghouse	Contract counterparty	Contract counterparty
Contract terms	Standardized	Negotiated	Negotiated
Settlement realization	Daily settlement	At expiration	Periodically (e.g., semiannual)

Forwards, futures, and swaps are the most common derivative contracts that represent a firm commitment for both counterparties.

A forward contract is an over-the-counter (OTC) derivative; two counterparties agree that the buyer will purchase an underlying from the seller at a prearranged fixed price. As OTC derivatives, forwards offer greater flexibility and customization than exchange-traded derivatives, but they usually involve more counterparty risk. Forwards are useful for importers and exporters to manage the currency risk in their accounts receivable and accounts payable.

At delivery, the payoff for the buyer is $S_T - F_0(T)$, and the payoff for the seller is $F_0(T) - S_T$.

- If $S_T > F_0(T)$, then the buyer has a positive payoff since she is paying less for the asset than its price on that date.
- If $S_T < F_0(T)$, then the seller has a positive payoff since she has sold the asset for more than its price on that date.



Example 1 Forward Market Settlement

An investor enters a cash-settled forward contract to purchase 100 ounces of gold at a forward price ($F_0(T)$) of \$1,077 in 6 months. Today's spot price for gold is \$1,000/oz, and, as a result, the spot price (S_T) is \$1,090.

The payoff is $S_T - F_0(T)$, or $1090 - 1077 = \$13$ per ounce.

The investor would receive \$1,300 ($\13×100) from the financial intermediary.

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David Baud



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web the balanced chemical equation for this reaction is $\text{mg s} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg no}_3 \text{ 2 aq} + \text{h}_2 \text{ g}$ in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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web magnesium nitric acid magnesium nitrate dihydrogen $\text{mg hno}_3 \text{ mg no}_3 \text{ 2} + \text{h}_2$ is a single displacement substitution reaction where one mole of solid magnesium mg and two moles of aqueous nitric acid hno 3 react to form one mole of aqueous magnesium nitrate mg no 3 2 and one mole of dihydrogen h 2 gas show chemical

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web when balancing chemical equations our goal is to have the same number of each type of atom on both sides of the equation only change the coefficients these are the numbers in front substances never change the subscripts the small numbers after elements

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web mar 4 2018 $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg no}_3 \text{ 2} + \text{h}_2$ on the left side there are mg 1 magnesium in elemental form h 2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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web magnesium carbonate nitric acid magnesium nitrate water carbon dioxide one mole of magnesium carbonate mgco_3

and two moles of nitric acid HNO_3 react to form one mole of magnesium nitrate $\text{Mg(NO}_3)_2$ one mole of water H_2O and one mole of carbon dioxide CO_2 show chemical structure image reactants

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web magnesium is already balanced on the reactant and product side balancing nitrogen by multiplying HNO_3 by 2 we get $\text{Mg s magnesium 2 HNO}_3 \text{1 nitric acid Mg(NO}_3)_2 \text{s magnesium nitrate H}_2 \text{g hydrogen gas oxygen and hydrogen get automatically balanced by balancing nitrogen}$

$\text{Mg(NO}_3)_2 \text{H}_2\text{SO}_4 \text{MgSO}_4 \text{HNO}_3$ chemical equation - Apr 07 2023

web magnesium nitrate sulfuric acid magnesium sulfate nitric acid

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web the balanced net ionic equation for the reaction of magnesium carbonate with nitric acid is $\text{MgCO}_3 \text{s} + 2\text{H}^+ \text{aq} \rightarrow \text{Mg}^{2+} \text{aq} + \text{CO}_2 \text{g} + \text{H}_2\text{O} \text{l}$ b $\text{Mg}^{2+} \text{aq} + \text{CO}_3^{2-} \text{aq} + 2\text{H}^+ \text{aq} \rightarrow 2\text{NO}_3^- \text{aq}$ mg write the balanced equation when nitric acid reacts with calcium hydroxide to produce water and calcium nitrate

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web how to balance $\text{HNO}_3 \text{Mg Mg(NO}_3)_2 \text{H}_2$ word equation nitric acid magnesium magnesium nitrate hydrogen gas type of chemical reaction for this reaction we have a single displacement reaction balancing strategies in this single displacement reaction the mg replaces the h in the HNO_3

how to balance $\text{Mg HNO}_3 \text{Mg(NO}_3)_2 \text{NO}_2 \text{H}_2\text{O}$ magnesium youtube - May 08 2023

web nov 20 2020 in this video we ll balance the equation $\text{Mg HNO}_3 \text{Mg(NO}_3)_2 \text{NO}_2 \text{H}_2\text{O}$ and provide the correct coefficients for each compound to balance $\text{Mg HNO}_3 \text{Mg(NO}_3)_2 \text{NO}_2 \text{H}_2\text{O}$ you ll need to be

lecture 13 acid base reactions widener university - Dec 03 2022

web magnesium hydroxide nitric acid balanced total equation $\text{Mg(OH)}_2 \text{aq} + 2\text{HNO}_3 \text{aq} \rightarrow 2\text{H}_2\text{O} \text{mg(NO}_3)_2 \text{aq}$ total ionic equation $\text{Mg}^{2+} \text{2 OH}^- \text{2 H}^+ \text{2 NO}_3^- \rightarrow 2\text{H}_2\text{O} \text{Mg}^{2+} \text{2 NO}_3^-$ net ionic equation $2\text{OH}^- \text{2 H}^+ \rightarrow 2\text{H}_2\text{O}$ sulfuric acid and lithium hydroxide

balanced total equation $\text{H}_2\text{SO}_4 + 2\text{LiOH} \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$