

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question : **[7]**

(i) The Arrhenius equation is _____.

(a) $k = A e^{\frac{RT}{E_a}}$

(b) $A = k e^{\frac{-E_a}{RT}}$

(c) $k = A e^{\frac{-RT}{E_a}}$

(d) $k = A e^{\frac{-E_a}{RT}}$

(ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol⁻¹, the entropy of vaporization will be _____.

(a) 4.0 J . K⁻¹ . mol⁻¹

(b) 3.0 J . K⁻¹ . mol⁻¹

(c) 1.5 J . K⁻¹ . mol⁻¹

(d) 0.5 J . K⁻¹ . mol⁻¹

(iii) The atomicity of sulphur in orthorhombic sulphur is _____.

(a) 8

(b) 6

(c) 4

(d) 2

(iv) The major binding force in diamond is _____.

(a) Covalent bond

(b) Ionic bond

(c) Metallic bond

(d) Co-ordinate covalent bond

(v) The boiling point of water at high altitude is low, because _____.

(a) the temperature is low.

(b) the atmospheric pressure is low.

(c) the temperature is high.

(d) the atmospheric pressure is high.

(vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is _____.

(a) 90 mhos.cm².

(b) 110 mhos.cm².mol⁻¹

(c) 200 mhos.cm².mol⁻¹

(d) 400 mhos.cm².mol⁻¹

(vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?

(a) Polling

(b) Pyrometallurgy

(c) Hydrometallurgy

(d) Calcination

Q. 2. Answer any THREE of the following : **[9]**

(i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –

(a) Magnetic property

(b) Oxidation state

(c) Hydrides

(ii) What is the value of ΔS_{sur} for the following reaction at 298 K –



Given that : $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.

(iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.

(iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

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Table of Contents Hsc Chemistry 2014

1. Understanding the eBook Hsc Chemistry 2014
 - The Rise of Digital Reading Hsc Chemistry 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Hsc Chemistry 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc Chemistry 2014
 - Personalized Recommendations
 - Hsc Chemistry 2014 User Reviews and Ratings
 - Hsc Chemistry 2014 and Bestseller Lists
5. Accessing Hsc Chemistry 2014 Free and Paid eBooks
 - Hsc Chemistry 2014 Public Domain eBooks
 - Hsc Chemistry 2014 eBook Subscription Services
 - Hsc Chemistry 2014 Budget-Friendly Options
6. Navigating Hsc Chemistry 2014 eBook Formats

- ePub, PDF, MOBI, and More
 - Hsc Chemistry 2014 Compatibility with Devices
 - Hsc Chemistry 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry 2014
 - Highlighting and Note-Taking Hsc Chemistry 2014
 - Interactive Elements Hsc Chemistry 2014
 8. Staying Engaged with Hsc Chemistry 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry 2014
 9. Balancing eBooks and Physical Books Hsc Chemistry 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry 2014
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Hsc Chemistry 2014
 - Setting Reading Goals Hsc Chemistry 2014
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Hsc Chemistry 2014
 - Fact-Checking eBook Content of Hsc Chemistry 2014
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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1638 1637 1636 1635 1634 1633 1632 1631 1630 1629 1628 1627 1626 1625 1624 1623 1622 1621 1620 1619 1618 1617 1616 1615 1614 1613 1612 1611 1610 1609 1608 1607 1606 1605 1604 1603 1602 1601 1600 1599 1598 1597 1596 1595 1594 1593 1592 1591 1590 1589 1588 1587 1586 1585 1584 1583 1582 1581 1580 1579 1578 1577 1576 1575 1574 1573 1572 1571 1570 1569 1568 1567 1566 1565 1564 1563 1562 1561 1560 1559 1558 1557 1556 1555 1554 1553 1552 1551 1550 1549 1548 1547 1546 1545 1544 1543 1542 1541 1540 1539 1538 1537 1536 1535 1534 1533 1532 1531 1530 1529 1528 1527 1526 1525 1524 1523 1522 1521 1520 1519 1518 1517 1516 1515 1514 1513 1512 1511 1510 1509 1508 1507 1506 1505 1504 1503 1502 1501 1500 1499 1498 1497 1496 1495 1494 1493 1492 1491 1490 1489 1488 1487 1486 1485 1484 1483 1482 1481 1480 1479 1478 1477 1476 1475 1474 1473 1472 1471 1470 1469 1468 1467 1466 1465 1464 1463 1462 1461 1460 1459 1458 1457 1456 1455 1454 1453 1452 1451 1450 1449 1448 1447 1446 1445 1444 1443 1442 1441 1440 1439 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1038 1037 1036 1035 1034 1033 1032 1031 1030 1029 1028 1027 1026 1025 1024 1023 1022 1021 1020 1019 1018 1017 1016 1015 1014 1013 1012 1011 1010 1009 1008 1007 1006 1005 1004 1003 1002 1001 1000 999 998 997 996 995 994 993 992 991 990 989 988 987 986 985 984 983 982 981 980 979 978 977 976 975 974 973 972 971 970 969 968 967 966 965 964 963 962 961 960 959 958 957 956 955 954 953 952 951 950 949 948 947 946 945 944 943 942 941 940 939 938 937 936 935 934 933 932 931 930 929 928 927 926 925 924 923 922 921 920 919 918 917 916 915 914 913 912 911 910 909 908 907 906 905 904 903 902 901 900 899 898 897 896 895 894 893 892 891 890 889 888 887 886 885 884 883 882 881 880 879 878 877 876 875 874 873 872 871 870 869 868 867 866 865 864 863 862 861 860 859 858 857 856 855 854 853 852 851 850 849 848 847 846 845 844 843 842 841 840 839 838 837 836 835 834 833 832 831 830 829 828 827 826 825 824 823 822 821 820 819 818 817 816 815 814 813 812 811 810 809 808 807 806 805 804 803 802 801 800 799 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