

26. The number of significant figures in the measured value 0.0204 is
 (1) Five (2) Three (3) Four (4) Two

Sol. Answer (2)

The non-zero digits after the decimal places are significant.

27. The number of significant figures in the measured value 26000 is
 (1) Five (2) Two (3) Three (4) Infinite

Sol. Answer (2)

The trailing zeros are not significant.
 So, only two digits are significant.

28. The number of significant zeroes present in the measured value 0.020040, is
 (1) Five (2) Two (3) One (4) Three

Sol. Answer (4)

Zeros appearing between and after non-zero numbers are significant.

0.020040

29. The number of significant figures in the measured value 4.700 m is the same as that in the value
 (1) 4700 m (2) 0.047 m (3) 4070 m (4) 470.0 m

Sol. Answer (4)

4.700 \Rightarrow Four significant figures.

Also, 470.0 m \Rightarrow Four significant figures.

30. If a calculated value 2.7465 g contains only three significant figures, the two insignificant digits in it are
 (1) 2 and 7 (2) 7 and 4 (3) 6 and 5 (4) 4 and 6

Sol. Answer (3)

2.7465 g \Rightarrow Last two digits are most insignificant.

31. An object of mass 4.237 g occupies a volume 1.72 cm³. The density of the object to appropriate significant figures is

- (1) 2.46 g cm⁻³ (2) 2.463 g cm⁻³ (3) 2.5 g cm⁻³ (4) 2.50 g cm⁻³

Sol. Answer (1)

$$m = 4.237 \text{ g}$$

$$V = 1.72 \text{ cm}^3$$

$$\text{Density} = \frac{\text{Mass}}{\text{Volume}} = \frac{4.237 \text{ g}}{1.72 \text{ cm}^3}$$

$$\Rightarrow \boxed{d = 2.46 \text{ g cm}^{-3}}$$

32. Round off the value 2.845 to three significant figures.

- (1) 2.85 (2) 2.84 (3) 2.80 (4) 2.83

Sol. Answer (2)

$$2.845 \Rightarrow 2.84$$

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Emilie Sanchez

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