



# General Electric Lavarropas

**Mark Steyvers**



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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized  $\text{Ca(OH)}_2$  when calcium is put in water reacts as follows  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is related to  $\Delta G$  for a process through eqn 1 and as long as  $\Delta H$  and  $\Delta S$  are independent of temperature eqn 1 can be combined with eqn 2 to determine  $\Delta H$  and  $\Delta S$  from the temperature dependence of k

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is

used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice  
*k<sub>sp</sub> calcium hydroxide chem132 determination of the k<sub>sp</sub> of calcium* - Jun 24 2022

web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the  
 solution  $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^-$   $5 \times 10^{-2} \text{ M}$  calcium ions  $5 \times 10^{-2} \text{ M}$   $0 \text{ M}$  hydroxide ions calculate calcium hydroxide  $K_{sp}$  using the  
 concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the  $K_{sp}$  of a slightly  
 soluble basic salt and how solubility changes with temperature  $K_{sp}$  is another form of equilibrium expression that deals with  
 solubility the reaction that will take place is  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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 molarity 0.050 M volume of filtered  $\text{Ca(OH)}_2$  solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml OH<sup>-</sup> in saturated  $\text{Ca(OH)}_2$   
 in saturated  $\text{Ca(OH)}_2$  M  $K_{sp}$  of  $\text{Ca(OH)}_2$  average  $K_{sp}$  of  $\text{Ca(OH)}_2$

chemistry 1210 lab 11 experiment 11 determining the  $K_{sp}$  of calcium - Sep 08 2023

web the chemical equation used was  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product was determined to be  $K_{sp} \text{Ca(OH)}_2$  from the  
 titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium  
 hydroxide was filtered before starting the titration

18.1 solubility product constant  $K_{sp}$  chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3\text{PO}_4$  is a constant it does not appear explicitly in  
 the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  
 $K_{sp} \text{Ca}_3\text{PO}_4 = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]$

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web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated  $\text{Ca(OH)}_2$  solution with a standard  
 HCl solution and find the  $K_{sp}$  or the OH<sup>-</sup> concentration was found and then the concentration was found and u for this  
 experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the  $K_{sp}$  of calcium hydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of  
 calcium hydroxide with hydrochloric acid was investigated determining the  $K_{sp}$  of calcium hydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$   
 for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of  $\text{Ca(OH)}_2$  is represented in equation form lution of  $\text{Ca(OH)}_2$  ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for  $\text{Ca(OH)}_2$  was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the  $\text{Ca}^{2+}$  and  $\text{OH}^-$  concentrations in solution

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web calculate the ksp for calcium hydroxide  $\text{Ca(OH)}_2$   $K_{sp} = 5 \times 10^{-6}$  find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is  $5 \times 10^{-6}$

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