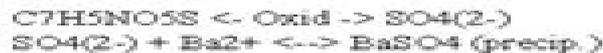


Problemas resueltos de Química analítica. PQA 4.20.

4.20. Una muestra de sidra de 100.00 g se somete a un proceso de oxidación mediante el cual su contenido en sacarina es transformado en sulfato, que posteriormente es precipitado con bario. Si la masa de este precipitado es de 0.0223 g, calcular el porcentaje de sacarina en la muestra de sidra

La sacarina es oxidada a sulfato y el sulfato es precipitado con bario según las reacciones siguientes y relaciones estequiométricas:



$$\text{moles sacarina/moles S} = \text{moles SO}_4^{2-} = \text{moles BaSO}_4$$

$$\text{moles sacarina} = \text{g BaSO}_4 / \text{Pm BaSO}_4 = 0.0223 / 233.40 = 9.55 \cdot 10^{-5}$$

$$\text{g sacarina} = 9.55 \cdot 10^{-5} \cdot \text{Pm sacarina} = 9.55 \cdot 10^{-5} \cdot 183.19 = 0.0175 \text{ g}$$

$$\% \text{ sacarina} = 0.0175 \cdot 100 / 100.00 = 0.0175\%$$

Problemas resueltos de Química analítica. PQA 4.19.

4.19. Una muestra de 1.227 g de alumbre, $\text{K}_2\text{SO}_4 \cdot \text{Al}_2(\text{SO}_4)_3 \cdot 24 \text{H}_2\text{O}$, que contiene sólo impurezas inertes, se disuelve y el aluminio se precipita con 8-hidroxiquinoleína. Una vez filtrado, el precipitado se calcina a Al_2O_3 , obteniéndose una masa de 0.1098 g. Calcular los porcentajes de aluminio y azufre en la muestra.

El aluminio precipita con la 8-hidroxiquinoleína $\text{C}_9\text{H}_7\text{NO}$, según la reacción:



Una vez calcinado a Al_2O_3 y determinado su peso, 0.1098 g, se puede calcular el contenido de aluminio en la muestra.

$$\begin{aligned} \text{g Al} &= \text{g Al}_2\text{O}_3 \cdot 2 \cdot \text{Pat Al} / \text{Pm Al}_2\text{O}_3 = \\ &= 0.1098 \cdot 2 \cdot 26.98 / 101.96 = 0.05811 \text{ g} \end{aligned}$$

$$\% \text{ Al} = 0.05811 \cdot 100 / 1.2277 = 4.73\%$$

Para calcular el contenido de azufre en peso y su porcentaje tendremos en cuenta la reacción estequiométrica:

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Ying-Ying Zheng



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Table of Contents 66 Problemas Resueltos De Ocwus

1. Understanding the eBook 66 Problemas Resueltos De Ocwus
 - The Rise of Digital Reading 66 Problemas Resueltos De Ocwus
 - Advantages of eBooks Over Traditional Books
2. Identifying 66 Problemas Resueltos De Ocwus
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 66 Problemas Resueltos De Ocwus
 - User-Friendly Interface
4. Exploring eBook Recommendations from 66 Problemas Resueltos De Ocwus
 - Personalized Recommendations
 - 66 Problemas Resueltos De Ocwus User Reviews and Ratings
 - 66 Problemas Resueltos De Ocwus and Bestseller Lists
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- 66 Problemas Resueltos De Ocwus Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - 66 Problemas Resueltos De Ocwus Compatibility with Devices
 - 66 Problemas Resueltos De Ocwus Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 66 Problemas Resueltos De Ocwus
 - Highlighting and Note-Taking 66 Problemas Resueltos De Ocwus
 - Interactive Elements 66 Problemas Resueltos De Ocwus
- 8. Staying Engaged with 66 Problemas Resueltos De Ocwus
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 66 Problemas Resueltos De Ocwus
- 9. Balancing eBooks and Physical Books 66 Problemas Resueltos De Ocwus
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 66 Problemas Resueltos De Ocwus
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 66 Problemas Resueltos De Ocwus
 - Setting Reading Goals 66 Problemas Resueltos De Ocwus
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 66 Problemas Resueltos De Ocwus
 - Fact-Checking eBook Content of 66 Problemas Resueltos De Ocwus
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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