9. ^{125}I has a half-life of 8.04 days. Assuming you start with a 1.35 mg sample of ^{125}I, how much will remain after 13.0 days?

   A) 0.825 mg
   B) 0.260 mg
   C) 0.622 mg
   D) 0.440 mg

$$k = \frac{\ln 2}{t_{1/2}} = \frac{\ln 2}{8.04 \text{ days}} = 0.0862 \text{ days}^{-1}$$

$$\ln N_t = -kt + \ln N_0$$

$$\ln N_t = -(0.0862 \text{ days}^{-1})(13.0 \text{ days}) + \ln(1.35 \text{ mg}) = -0.821$$

$$N_t = e^{-0.821} = \textbf{0.440 mg}$$

10. Which one of the following devices converts radioactive emissions to light for detection?

   A) Geiger counter
   B) photographic film
   C) scintillation counter
   D) none of these

11. Which one of the following is true?

   A) Most spontaneous nuclear reactions are exothermic.
   B) Some spontaneous nuclear reactions are endothermic.
   C) All spontaneous nuclear reactions are exothermic.
   D) There is no relationship between exothermicity and spontaneity in nuclear reactions.

12. Binding energy per nucleon is greatest for nuclei with...

   A) large mass numbers.
   B) intermediate mass numbers.
   C) small mass numbers.

13. When two atoms of ^2H are fused to form one atom of ^4He, the total energy evolved is 3.863×10^{-12} J. What is the total change in mass for this reaction? ($c = 3.00 \times 10^8$ m/s)

   A) 1.29×10^{-21} g
   B) 4.29×10^{-29} g
   C) 1.29×10^{-29} g
   D) 4.29×10^{-21} g

$$E = mc^2 \quad \rightarrow \quad m = \frac{E}{c^2} = \frac{3.863 \times 10^{-12} \text{ J}}{(3.00 \times 10^8 \text{ m/s})^2} = 4.29 \times 10^{-29} \text{ kg} \times \left(\frac{1000 \text{ g}}{1 \text{ kg}}\right) = \textbf{4.29} \times 10^{-26} \text{ g}$$

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