



### PRE-LAB QUESTIONS

1. What would you estimate the pH of hydrochloric acid, a strong acid, to be? What is the pH of water? The pH of HCl is 3.01, so yes a strong acid, the pH of water is 7
2. Use the Henderson-Hasselbalch equation to estimate the pH of a buffer solution that is composed of 20 mL of 10 M sodium formate and 20 mL of 1 M formic acid (Note: the pKa of formic acid is 3.75). 4.75 would be the pH
3. Look at Table 4 in the procedure portion of the experiment. Calculate the pH you would expect each of the buffer solutions (A, B, C, D, and E) to be using the Henderson-Hasselbalch equation, assuming that the solutions of acetic acid and sodium acetate are equimolar.

# Properties Of Buffer Solutions Pre Lab Answers

**Scott C. Dulebohn**



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