

### Preparation of a standard solution

**Aim:** To prepare  $2.0 \text{ mol dm}^{-3}$  standard sodium chloride solution using a  $250 \text{ cm}^3$  volumetric flask.

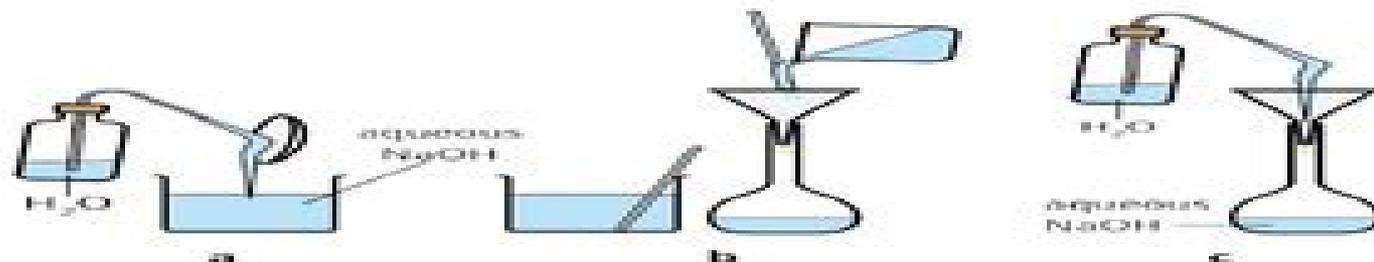
**Tools 1:** Experimental techniques – Applying techniques.

**Available Apparatus:** digital balance, volumetric flask, pipettes, and pipette bulbs, spatula, watch glass, beaker, funnel, stirring rod, and wash bottle.

**Reagents:** Sodium chloride, food coloring.

#### Preparation of sodium hydroxide solution.

1. Calculate the mass (in grams) of solid sodium hydroxide required to prepare  $2.0 \text{ mol dm}^{-3}$  solution in the assigned volumetric flask.
2. Place a watch glass on the analytical balance and press the 'Tare' button.
3. Accurately weigh approximately the amount of sodium chloride, you have calculated you need, recording the exact amount used.
4. Add approximately  $50 \text{ cm}^3$  distilled water to the beaker and stir with a glass rod until the solid has dissolved.
5. Using a funnel, carefully transfer your solution to the volumetric flask given.
6. Add approximately  $10 \text{ cm}^3$  distilled water to the beaker, rinse it around the beaker and add to the volumetric flask. Repeat with a further  $10 \text{ cm}^3$  portion.
7. Carefully add distilled water to the volumetric flask, running it around the edge of the funnel to wash any remaining NaCl in, and also running it over the stirring rod to wash this in.
8. When the level of solution is within  $1 \text{ cm}$  of the graduation on the flask, add distilled water drop-wise using a pipette until the bottom of the meniscus is in line with the graduation mark.
9. Firmly place the lid on the volumetric flask, and invert it several times, holding the lid in place. You will know your solution is fully mixed when inverting it no longer makes it look 'swirly'.
10. Calculate the exact concentration (4 s.f.) of the solution, using the mass measured in Step 2 and label the flask with the solution, the concentration, the date of preparation and your name.
11. The solution should be stored in a cool dark place until it is to be used.



A. Calculate the mass of sodium chloride required to prepare a solution  $2.0 \text{ mol dm}^{-3}$  solution.

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# Preparation Of Solutions Lab Report

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## **Preparation Of Solutions Lab Report:**

**Laboratory Methods in Microfluidics** Basant Giri, 2017-05-15 Laboratory Methods in Microfluidics features a range of lab methods and techniques necessary to fully understand microfluidic technology applications Microfluidics deals with the manipulation of small volumes of fluids at sub millimeter scale domain channels This exciting new field is becoming an increasingly popular subject both for research and education in various disciplines of science including chemistry chemical engineering and environmental science The unique properties of microfluidic technologies such as rapid sample processing and precise control of fluids in assay have made them attractive candidates to replace traditional experimental approaches Practical for students instructors and researchers this book provides a much needed comprehensive new laboratory reference in this rapidly growing and exciting new field of research Provides a number of detailed methods and instructions for experiments in microfluidics Features an appendix that highlights several standard laboratory techniques including reagent preparation plus a list of materials vendors for quick reference Authored by a microfluidics expert with nearly a decade of research on the subject

**Science Educator's Guide to Laboratory Assessment** Rodney L. Doran, 2002 The book opens with an up to date discussion of assessment theory research and uses Then comes a wealth of sample assessment activities in biology chemistry physics and Earth science Keyed to the National Science Education Standards the activities include reproducible task sheets and scoring rubrics All are ideal for helping students reflect on their own learning during science lab

**Laboratory Methods in Dynamic Electroanalysis** M. Teresa Fernández Abedul, 2019-10-13 Laboratory Methods in Dynamic Electroanalysis is a useful guide to introduce analytical chemists and scientists of related disciplines to the world of dynamic electroanalysis using simple and low cost methods The trend toward decentralization of analysis has made this fascinating field one of the fastest growing branches of analytical chemistry As electroanalytical devices have moved from conventional electrochemical cells 10 20 mL to current cells e g 5 50 mL based on different materials such as paper or polymers that integrate thick or thin film electrodes interesting strategies have emerged such as the combination of microfluidic cells and biosensing or nanostructuring of electrodes This book provides detailed easy procedures for dynamic electroanalysis and covers the main trends in electrochemical cells and electrodes including microfluidic electrodes electrochemical detection in microchip electrophoresis nanostructuring of electrodes development of bio enzymatic immuno and DNA assays paper based electrodes interdigitated array electrodes multiplexed analysis and combination with optics Different strategies and techniques amperometric voltammetric and impedimetric are presented in a didactic practice based way and a bibliography provides readers with additional sources of information Provides easy to implement experiments using low cost simple equipment Includes laboratory methodologies that utilize both conventional designs and the latest trends in dynamic electroanalysis Goes beyond the fundamentals covered in other books focusing instead on practical applications of electroanalysis

**Chemistry in the Laboratory** James M. Postma, Julian L. Robert, J. Leland

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**Working with Chemistry** Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

**Bibliography of Scientific and Industrial Reports**, 1947 Annual Report New York (State). Department of Health, 1925 Vols for 1949 issued in 2 vols New York s health and statistical part

**Laboratory Exercises for Preparatory Chemistry** Kathy Dodds Tyner, 1994-06 Laboratory Exercises for Preparatory Chemistry is the perfect complement to a one semester preparatory chemistry laboratory course Tyner s manual emphasizes the application of chemistry and the principles of science to everyday life The labs are directly applicable to the real world and often contain supplemental assignments that illustrate an application

Annual Reports of the Connecticut Agricultural Experiment Station Connecticut Agricultural Experiment Station, 1910

*Chemical Investigations* Nancy Konigsberg Kerner, 1986

**Annual Report of the Division of Laboratories and Research** New York (State). Department of Health. Division of Laboratories and Research, 1921

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**Ceramic Abstracts**, 1930 **Annual Report of the State Department of Health of New York for the Year Ending December 31 ...** New York (State). Department of Health, 1925

**American Journal of Public Health**, 1924 Includes section Books and reports

*The Extra Pharmacopœia of Martindale and Westcott* William Martindale, 1920

*Journal of the American Medical Association* American Medical Association, 1909 Includes proceedings of the Association papers read at the annual sessions and list of current medical literature

*Journal of the American Public Health Association*, 1924

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