



Wifi 7 Router Apple Watch Guide

Tobias Bleicker



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$2 naf$ $1 cl_2$ $2 h_2$ $1 o_2$ Æ $2 h_2o$ pb oh 2 $2 hcl$ Æ $2 h_2o$ $1 pbcl_2$ $albr_3$ $3 k_2so_4$ Æ $6 kbr$ $1 al_2 so_4$ $3 ch_4$ $2 o_2$ Æ $1 co_2$ $2 h_2o$ c_3h_8 5

o_2 Æ $3 co_2$ $4 h_2o$ c_8h_{18} $25 o_2$ Æ $16 co_2$ $18 h_2o$

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carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas gaseous butane $c_4 h_{10}$ reacts with

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are interpreted as containing the number 1 hen balancing $1 na$ $3 po$ $4 3 koh$ $3 naoh$ $1 k$ $3 po$ $4 1 mgf$ $2 1 li$ $2 co$ $3 1 mgco$ $3 2$

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 $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
 $2 \text{ NaCl} + \text{F}_2 \rightarrow 2 \text{ NaF} + \text{Cl}_2$
 $4 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$
 $5 \text{ Pb} + 2 \text{ OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 + 6 \text{ Al} + 3 \text{ K}_2\text{SO}_4 + 4 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 + 3 \text{ CH}_4 + 2 \text{ O}_2 + 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

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 $1 \text{ O}_2 + 3 \text{ NH}_3 \rightarrow \text{HNO}_3 + \text{H}_2\text{O} + 2 \text{ O}_2 + 3 \text{ H}_2\text{O} + 2 \text{ H}_2\text{O} + 2 \text{ RbH}_2 + \text{O} + \text{RbOH} + \text{H}_2 + 5 \text{ Cl}_2 + 2 \text{ KBr} + \text{KCl} + \text{Br}_2 + 6 \text{ N}_2 + 2 \text{ H}_2 + \text{NH}_3 + 7 \text{ N}_2\text{F}_6 + 3 \text{ Ca} + \text{Nd} + \text{CaF}_2 + 8 \text{ CH}_4 + \text{H}_2 + \text{O} + \text{CO} + \text{H}_2 + 9 \text{ C} + 2 \text{ H}_2 + 5 \text{ OH}^-$

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 $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
 $2 \text{ NaCl} + \text{F}_2 \rightarrow 2 \text{ NaF} + \text{Cl}_2$
 $4 \text{ H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$
 $5 \text{ Pb} + 2 \text{ OH}^- + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2 + 6 \text{ Al} + 3 \text{ K}_2\text{SO}_4 + 4 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 + 3 \text{ CH}_4 + 2 \text{ O}_2 + 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

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coefficients equal to one 1 do not need to be shown in your answers
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 $4 \text{ Fe} + 3 \text{ O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3$
 $2 \text{ Fe}_2\text{O}_3 + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe}_2\text{SO}_4 + 3 \text{ H}_2\text{O} + 3 \text{ H}_2\text{O} + 2 \text{ C}_2\text{H}_4\text{O}_2$

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 $38 \text{ Fe} + 2 \text{ O}_3 + 3 \text{ CO} + 2 \text{ Fe}_3\text{CO}_2 + 39 \text{ H}_2 + 3 \text{ PO}_4 + 3 \text{ BaO} + 3 \text{ PO}_4 + 2 \text{ H}_2 + \text{O} + 40 \text{ Mg} + \text{NH}_4 + \text{PO}_4 + \text{Mg} + 2 \text{ P}_2 + \text{O}_7 + 2 \text{ NH}_3 + \text{H}_2 + \text{O} + 41 \text{ Al} + 3 \text{ H}_2 + \text{SO}_4 + \text{Al}_2 + \text{SO}_4 + 3 \text{ H}_2 + 42 \text{ P}_4 + \text{O} + 10 \text{ HCl} + \text{POCl}_3 + 3 \text{ HPO}_3 + 43 \text{ HNO}_2 + \text{H}_2 + \text{O} + 30 + 2 \text{ HNO}_2$

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