

## Question 35



In the early 1900s, Robert Millikan performed a series of experiments to measure the charge on tiny oil droplets suspended in an electric field. By analyzing how the droplets moved under gravity and an applied voltage, Millikan determined the fundamental electric charge (the charge on a single electron). From this charge and additional data, the mass of the electron could be calculated. In this passage, you will examine data from similar oil drop experiments to investigate the mass of the electron.



A fine mist of oil droplets is introduced into a transparent chamber situated between two horizontal metal plates. The space between the plates is observed through a microscope or viewing scope. A DC power supply applies a variable voltage across the plates, creating a uniform electric field. The droplets become charged locally through friction or ionization and can be made to rise, fall more slowly, or remain suspended depending on the balance of forces on them. The time it takes for droplets to move between reference markers and the voltage across the plates are recorded. A cross-sectional view: two parallel horizontal plates, with an oil droplet in between. A microscope is shown to the side for observing the droplet's motion. The top

In Experiment 1, why is the droplet's time to fall measured with the electric field turned off?

- A. To observe how the droplet interacts with the charged plates
- B. To determine the droplet's mass-related properties free from electrical forces
- C. To calculate the exact charge on the droplet
- D. To allow the droplet to pick up additional charge

# Act Practice Top Setup

**Joan U. Levy, Norman Levy**



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### **Table of Contents Act Practice Top Setup**

1. Understanding the eBook Act Practice Top Setup
  - The Rise of Digital Reading Act Practice Top Setup
  - Advantages of eBooks Over Traditional Books
2. Identifying Act Practice Top Setup
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Act Practice Top Setup
  - User-Friendly Interface
4. Exploring eBook Recommendations from Act Practice Top Setup
  - Personalized Recommendations
  - Act Practice Top Setup User Reviews and Ratings
  - Act Practice Top Setup and Bestseller Lists
5. Accessing Act Practice Top Setup Free and Paid eBooks

- Act Practice Top Setup Public Domain eBooks
  - Act Practice Top Setup eBook Subscription Services
  - Act Practice Top Setup Budget-Friendly Options
6. Navigating Act Practice Top Setup eBook Formats
- ePub, PDF, MOBI, and More
  - Act Practice Top Setup Compatibility with Devices
  - Act Practice Top Setup Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Act Practice Top Setup
  - Highlighting and Note-Taking Act Practice Top Setup
  - Interactive Elements Act Practice Top Setup
8. Staying Engaged with Act Practice Top Setup
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Act Practice Top Setup
9. Balancing eBooks and Physical Books Act Practice Top Setup
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Act Practice Top Setup
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Act Practice Top Setup
- Setting Reading Goals Act Practice Top Setup
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Act Practice Top Setup
- Fact-Checking eBook Content of Act Practice Top Setup
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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