

## AZEOTROPIC DATA FOR BINARY MIXTURES

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Binary homogeneous (single-phase) liquid mixtures having an extremum (maximum or minimum) vapor pressure  $P$  at constant temperature  $T$ , as a function of composition, are called azeotropic mixtures, or simply azeotropes. The composition is usually expressed as mole fractions, where  $x_i$  for component  $i$  in the liquid phase and  $y_i$  for component  $i$  in the vapor phase are identical. Mixtures that do not show a maximum or minimum are called zeotropic. A maximum (minimum) of the  $P(x_i)$  or  $P(y_i)$  curves corresponds to a minimum (maximum) of the boiling temperature  $T$  at constant  $P$  plotted as a function of  $x_i$  or  $y_i$ , [see  $T(x_i)$  and  $T(y_i)$  curves, Types I and II, in Fig. 1]. Azeotropes in which the pressure is a maximum (temperature is a minimum) are often called positive azeotropes, while pressure-minimum (temperature-maximum) azeotropes are called negative azeotropes. The coordinates of an azeotropic point are the azeotropic temperature  $T_{az}$ , pressure  $P_{az}$ , and the vapor-phase composition  $y_{1,az}$ , which is the same as the liquid-phase composition  $x_{1,az}$ .

In the two-phase liquid-liquid region of partially miscible (heterogeneous) mixtures, the vapor pressure at constant  $T$  (or the boiling temperature at constant  $P$ ) is independent of the global composition  $x_i$  of the two coexisting liquid phases between the equilibrium compositions  $x_i^*$  and  $x_i^{\prime}$  ( $x_i^* < x_i^{\prime}$ )

ration is impossible by ordinary distillation, even with an infinitely large number of stages.

Following eq. (1) azeotropic behavior will always occur in homogeneous binary systems when the vapor pressure ratio  $P_1^s/P_2^s$  is equal to the ratio of the activity coefficients  $\gamma_1/\gamma_2$ .

Various thermodynamic methods based on  $g^E$ -models (Wilson, NRTL, UNIQUAC) or group contribution methods (UNIFAC, modified UNIFAC, ASOG, PSRK) can be used for either calculating or predicting the required activity coefficients for the components under given conditions of temperature and composition (Reference 2).

Because of the importance of azeotropic data for the design of distillation processes, compilations have been available in book form for quite some time (References 3-7). The most recent printed data collection was published in 1994 (Reference 8). A revised and extended version appeared in 2004 (Reference 9).

A collection of approximately 47,400 zeotropic and azeotropic data sets, compiled from 6600 references, are stored in a comprehensive computerized data bank (Reference 10). The references from the above-mentioned compilations and from the vapor-liquid equilibrium part of the Dortmund Data Bank (Reference 11) were supplemented by references found from CAS online

# Azeotropic Data For Binary Mixtures

**K Morrison**



## **Azeotropic Data For Binary Mixtures :**

The book delves into Azeotropic Data For Binary Mixtures . Azeotropic Data For Binary Mixtures is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Azeotropic Data For Binary Mixtures , encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Azeotropic Data For Binary Mixtures
- Chapter 2: Essential Elements of Azeotropic Data For Binary Mixtures
- Chapter 3: Azeotropic Data For Binary Mixtures in Everyday Life
- Chapter 4: Azeotropic Data For Binary Mixtures in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Azeotropic Data For Binary Mixtures . This chapter will explore what Azeotropic Data For Binary Mixtures is, why Azeotropic Data For Binary Mixtures is vital, and how to effectively learn about Azeotropic Data For Binary Mixtures .
3. In chapter 2, the author will delve into the foundational concepts of Azeotropic Data For Binary Mixtures . This chapter will elucidate the essential principles that must be understood to grasp Azeotropic Data For Binary Mixtures in its entirety.
4. In chapter 3, this book will examine the practical applications of Azeotropic Data For Binary Mixtures in daily life. The third chapter will showcase real-world examples of how Azeotropic Data For Binary Mixtures can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Azeotropic Data For Binary Mixtures in specific contexts. The fourth chapter will explore how Azeotropic Data For Binary Mixtures is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Azeotropic Data For Binary Mixtures . This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Azeotropic Data For Binary Mixtures .

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