

Chapter 3 Chemistry Test: The Atom

Multiple Choice

Choose the choice that best completes the statement or answers the question.

- ____ 1. The law of conservation of mass follows from the concept that
- atoms are indestructible.
 - atoms of different elements have different properties.
 - matter is composed of atoms.
 - atoms can be destroyed in chemical reactions.
- ____ 2. Who explained the behavior of positively charged particles being deflected from a metal foil in the nucleus?
- Ernest Rutherford
 - John Dalton
 - James Chadwick
 - Walter Rastar
- ____ 3. In the gold foil experiment, most of the particles fired at the foil
- bounced back.
 - passed through the foil.
 - were absorbed by the foil.
 - combined with the foil.
- ____ 4. What did Rutherford conclude about the structure of the atom?
- An atom is indivisible.
 - Electrons make up the center of an atom.
 - An atom carries a positive charge.
 - An atom contains a small, dense, positively charged central region.
- ____ 5. A nuclear particle that has about the same mass as a proton, but with no electrical charge, is called a(n)
- neutron.
 - electron.
 - isotope.
 - proton.
- ____ 6. Which part of an atom has a mass approximately equal to $1/1836$ of the mass of a common hydrogen atom?
- neutron
 - electron
 - proton
 - electron cloud
- ____ 7. The mass of a neutron is
- about the same as that of a proton.
 - about the same as that of an electron.
 - double that of a proton.
 - double that of an electron.
- ____ 8. The nucleus of most atoms is composed of
- tightly packed protons.
 - tightly packed neutrons.
 - tightly packed protons and neutrons.
 - loosely connected protons and electrons.
- ____ 9. An aluminum isotope consists of 13 protons, 13 electrons, and 14 neutrons. Its mass number is
- 13.
 - 14.
 - 27.
 - 40.
- ____ 10. Isotopes are atoms of the same element that have different
- physical chemical properties.
 - masses.
 - numbers of protons.
 - numbers of electrons.
- ____ 11. Atoms of the same element that have different masses are called
- isotopes.
 - isotones.
 - isobars.
 - isomers.
- ____ 12. Most of the volume of an atom is occupied by the
- nucleus.
 - neutrons.
 - electron cloud.
 - protons.

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