



Meal Prep Ideas Top

Ava Manning



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{Au}^{3+} + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow \text{O}_2 + 2\text{H}_2\text{O} + 2\text{e}^-$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{Au}^{3+} + 4\text{e}^-$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 + \text{O}_2 + 2\text{H}_2\text{O} + 2\text{e}^- \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^- + 2\text{e}^-$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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