

AP Chemistry Chemical Kinetics Worksheet Answers

1.	Half lives	Reactant Fraction Remaining
	1	$\frac{1}{2}$
	2	$\frac{1}{4}$
	3	$\frac{1}{8}$
	4	$\frac{1}{16}$
	5	$\frac{1}{32}$

2. 2nd Order; 1st Order

3. a. True b. True

c. False; As [Reactant] decreases the rate will decrease d. True

4. a. $-1/2 \Delta[\text{O}_3]/\Delta t = 1/3 \Delta[\text{O}_2]/\Delta t$
b. $-1/2 \Delta[\text{HOF}]/\Delta t = 1/2 \Delta[\text{HF}]/\Delta t = \Delta[\text{O}_2]/\Delta t$
c. $-\Delta[\text{N}_2]/\Delta t = -1/3 \Delta[\text{H}_2]/\Delta t = 1/2 \Delta[\text{NH}_3]/\Delta t$

5. a. rate = $k[\text{NO}_2]^2[\text{CO}]^0$ or $k[\text{NO}_2]^2$
b. rate = 1/4 the original value

6. The initial rate $R = k[\text{B}]_0^2$

7. a. $6.0 \times 10^{-3} / 1.2 \times 10^{-2} = (0.10/0.20)^X$
 $X = 1$ so it is 1st order with respect to A
 $1.1 \times 10^{-1} / 1.2 \times 10^{-2} = (0.15/0.05)^X$
 $X = 2$ so it is 2nd order with respect to B
Rate Law = $k[\text{A}][\text{B}]^2$

b. Overall it is 3rd order

c. Probably not a single step since a 3-body collision would have to occur.

8. a. half life = $\ln(1/2) / -5.0 \times 10^{-4} \text{ s}^{-1} = 1.4 \times 10^3 \text{ s}$
 $-\ln(0.5) / 5.0 \times 10^{-4} \text{ s}^{-1} = 1.4 \times 10^3 \text{ s}$
b. Assign an original value of 1.00M
 $\ln(0.1/1.0) = -5.0 \times 10^{-4} \text{ s}^{-1} t$
 $t = -\ln(0.1) / 5.0 \times 10^{-4} \text{ s}^{-1} = 4.6 \times 10^3 \text{ s}$

9. After 30.min (one half-life)

$$P_{\text{HOF}} = \frac{1}{2} (100. \text{mmHg}) = 50. \text{mmHg}$$

$$P_{\text{HF}} = 50. \text{mmHg}$$

$$P_{\text{O}_2} = 25. \text{mmHg}$$

Calculating the rate constant k ,

$$\ln([\text{HOF}]/[\text{HOF}]_0) = -kt \qquad \ln(0.5) = -k(30. \text{min})$$

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Princeton Review



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