

Name _____ Date _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data for table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	2 molecules	2 molecules	NO	2 molecules O_2
2 molecules	3 molecules	3 molecules	O_2	3 molecules NO
1.00 mol	2.00 mol	2.00 mol	none	none
2.00 mol	2.00 mol	2.00 mol	NO	2.00 mol O_2
3.00 mol	3.00 mol	3.00 mol	O_2	3.00 mol NO
1.00 mol	3.00 mol	2.00 mol	NO	2.00 mol O_2
2.00 mol	3.00 mol	3.00 mol	O_2	1.00 mol NO
3.00 mol	3.00 mol	3.00 mol	NO	0.00 mol O_2
32.00 g	60.02 g	92.02 g	none	none
60.02 g	60.02 g	60.02 g	O_2	32.00 g NO
92.02 g	92.02 g	92.02 g	O_2	0.00 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

60.02 g NO = 1 mol NO /30.01 g NO = 0.004 mol NO

actual mole ratio = 0.004 mol NO /0.3125 mol O_2 = 2.122 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO / O_2 is the limiting reactant.

Mass of NO used = 0.3125 mol O_2 = 2 mol NO /1 mol O_2 = 0.6250 mol NO

0.6250 mol NO = 30.01 g NO /1 mol NO = 18.75 g NO

Mass of NO_2 produced = 0.6250 mol NO = 46.01 g NO_2 /1 mol NO = 28.75 g NO_2

Excess NO = 20.00 g NO - 18.75 g NO = 1.24 g NO

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CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the nitride from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and GaN are 69.72 g/mol and 122.62 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield volume = 7.42 g GaN
- mass of reactant = 7.00 g Ga
- number of moles of reactant = 7.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga/69.72 g Ga
- number of moles of product = 0.00 mol Ga \times $\frac{2 \text{ mol GaN}}{1 \text{ mol Ga}}$ = $\frac{0.100 \text{ mol Ga}}{2 \text{ mol Ga}}$ = 0.05 mol Ga
- theoretical yield = 0.0000 mol GaN = $\frac{122.62 \text{ g GaN}}{1 \text{ mol GaN}}$ correct
- percent yield = $\frac{7.42 \text{ g GaN}}{122.62 \text{ g GaN}} \times 100$ = $\frac{7.42 \text{ g GaN}}{6.87 \text{ g GaN}} \times 100$ = 106

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AW Rasmussen



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