

A weak base is titrated by a base. At the equivalence and beyond the base will completely remove the protons from the weak base of an indicator. Therefore, the color associated with the indicator will appear the color of the product of the reaction and not the color of the weak base or acid.



Example:

What is the pH of a 0.10 M solution of ammonia ($K_b = 1.8 \times 10^{-5}$) titrated with 0.10 M hydrochloric acid ($K_a = 1.0 \times 10^{-7}$) at the equivalence point?

Step 1: Write the reaction and determine the stoichiometry.

$$\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4^+ + \text{Cl}^-$$

Step 2: Determine the concentration of the weak acid at the equivalence point.

At the equivalence point, the concentration of the weak acid is 0.05 M.

Step 3: Calculate the pH of the weak acid solution.

The K_a of NH_4^+ is 5.6×10^{-10} .

$$\text{NH}_4^+ \rightleftharpoons \text{NH}_3 + \text{H}^+$$

Initial: 0.05 M, 0 M, 0 M
 Change: -x, +x, +x
 Equilibrium: 0.05 - x, x, x

$$K_a = \frac{[\text{NH}_3][\text{H}^+]}{[\text{NH}_4^+]} = \frac{x^2}{0.05 - x} = 5.6 \times 10^{-10}$$

$x = [\text{H}^+] = 5.3 \times 10^{-6} \text{ M}$

$\text{pH} = -\log(5.3 \times 10^{-6}) = 5.28$

- Answer the following questions. Show all work and report answers with units.
- Calculate the pH of a 0.10 M solution of ammonia ($K_b = 1.8 \times 10^{-5}$) titrated with 0.10 M hydrochloric acid ($K_a = 1.0 \times 10^{-7}$) at the equivalence point.
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Acid Base Titrations Chem Worksheet 19 5 Answers

James Sherwood Fritz



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