

9/2/16

S. Sai Krishna

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ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$\boxed{Q = \Delta E + P\Delta V}$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$\boxed{H = \Delta E + P\Delta V}$$

$$\boxed{H = \Delta E + \Delta n RT}$$

$\Delta n =$ no. of moles of Prod
- no. of moles of react.

→ For ΔH .

- If ΔH is -ve. reaction is exothermic.
- If ΔH is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is -322.1 kJ . Calculate the heat of combustion at constant pressure.



Engineering Chemistry 2 Notes

BM King



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Table of Contents Engineering Chemistry 2 Notes

1. Understanding the eBook Engineering Chemistry 2 Notes
 - The Rise of Digital Reading Engineering Chemistry 2 Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Engineering Chemistry 2 Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Engineering Chemistry 2 Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Engineering Chemistry 2 Notes
 - Personalized Recommendations

- Engineering Chemistry 2 Notes User Reviews and Ratings
- Engineering Chemistry 2 Notes and Bestseller Lists
- 5. Accessing Engineering Chemistry 2 Notes Free and Paid eBooks
 - Engineering Chemistry 2 Notes Public Domain eBooks
 - Engineering Chemistry 2 Notes eBook Subscription Services
 - Engineering Chemistry 2 Notes Budget-Friendly Options
- 6. Navigating Engineering Chemistry 2 Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Engineering Chemistry 2 Notes Compatibility with Devices
 - Engineering Chemistry 2 Notes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Engineering Chemistry 2 Notes
 - Highlighting and Note-Taking Engineering Chemistry 2 Notes
 - Interactive Elements Engineering Chemistry 2 Notes
- 8. Staying Engaged with Engineering Chemistry 2 Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Engineering Chemistry 2 Notes
- 9. Balancing eBooks and Physical Books Engineering Chemistry 2 Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Engineering Chemistry 2 Notes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Engineering Chemistry 2 Notes
 - Setting Reading Goals Engineering Chemistry 2 Notes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Engineering Chemistry 2 Notes
 - Fact-Checking eBook Content of Engineering Chemistry 2 Notes

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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