

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)**1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance** Basic SI Unit? **kg**
 - Volume of an object? **ruler or water displacement method** Basic SI Unit? **cm³ or mL**
 - Length of an object? **Ruler** Basic SI Unit? **meter**
 - Density of an object? **graduated cylinder** Basic SI Unit? **mL**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 33 g and has a volume of 3 cm³? **10g/cm³**
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. **solid H₂O less dense than liquid H₂O**

2. Chemistry Basics

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move fixed orbits around the atom's nucleus in specific layers, or electron shells . Atoms absorb or give off energy when the electrons move from one shell (energy level) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

4. Periodic Table

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
 - The atomic number represents the number of **protons**.
 - The atomic mass represents the number of **protons and neutrons**.

Physical Science Final Exam 3 Study Guide

Ying Liu



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Enjoying the Beat of Expression: An Mental Symphony within **Physical Science Final Exam 3 Study Guide**

In a world eaten by monitors and the ceaseless chatter of fast conversation, the melodic beauty and mental symphony developed by the written term usually fade into the back ground, eclipsed by the constant sound and disruptions that permeate our lives. However, located within the pages of **Physical Science Final Exam 3 Study Guide** a marvelous fictional value filled with fresh thoughts, lies an immersive symphony waiting to be embraced. Crafted by a wonderful composer of language, that charming masterpiece conducts readers on a psychological trip, well unraveling the hidden songs and profound affect resonating within each cautiously constructed phrase. Within the depths of this poignant assessment, we will examine the book is main harmonies, analyze their enthralling publishing fashion, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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