



PIPE FITTING FRICTION CALCULATION

The friction loss for fittings depends on a K factor which can be found in many sources such as the Crane's Hydraulic data book or the Hydraulic Institute Engineering data book, the charts which it represents are in Figures 1 and 2.

The fitting friction ΔH_{fit} can be calculated based on the following formula where K is a factor based on the type of fitting, v is the velocity in feet/second, g is the acceleration due to gravity (32.17 ft/s^2).

$$\Delta H_{fit} = K \cdot \frac{v^3}{2g \cdot D^5}$$

For example a 2" NPS iron screwed elbow has a K factor of 0.85 according to Figure 1 and using a velocity of 10 ft/s (this is determined from the flow rate). The fitting friction loss will be:

$$\Delta H_{fit} = K \cdot \frac{v^3}{2g \cdot D^5} = 0.85 \cdot \frac{10^3}{2 \times 32.17 \times 2^5} = 1.3$$

Pipe Fitting Friction Calculation Can Be Calculated Based

**American Society of Heating and Air-
Conditioning Engineers**



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Ludwig's Applied Process Design for Chemical and Petrochemical Plants A. Kayode Coker, 2011-08-30 This complete revision of *Applied Process Design for Chemical and Petrochemical Plants Volume 1* builds upon Ernest E Ludwig's classic text to further enhance its use as a chemical engineering process design manual of methods and proven fundamentals This new edition includes important supplemental mechanical and related data nomographs and charts Also included within are improved techniques and fundamental methodologies to guide the engineer in designing process equipment and applying chemical processes to properly detailed equipment All three volumes of *Applied Process Design for Chemical and Petrochemical Plants* serve the practicing engineer by providing organized design procedures details on the equipment suitable for application selection and charts in readily usable form Process engineers designers and operators will find more chemical petrochemical plant design data in *Volume 2 Third Edition* which covers distillation and packed towers as well as material on azeotropes and ideal non ideal systems *Volume 3 Third Edition* which covers heat transfer refrigeration systems compression surge drums and mechanical drivers A Kayode Coker is Chairman of Chemical Process Engineering Technology department at Jubail Industrial College in Saudi Arabia He is both a chartered scientist and a chartered chemical engineer for more than 15 years and an author of *Fortran Programs for Chemical Process Design Analysis and Simulation* Gulf Publishing Co and *Modeling of Chemical Kinetics and Reactor Design* Butterworth Heinemann Provides improved design manuals for methods and proven fundamentals of process design with related data and charts Covers a complete range of basic day to day petrochemical operation topics with new material on significant industry changes since 1995

Agriculture Handbook, 1949 Set includes revised editions of some issues

[Head Loss in Quick-coupled Aluminum Pipe Used for Sprinkler Irrigation Systems](#) William Oscar Ree, 1959

Building Technology Benjamin Stein, 1996 The complete guide to building technology This comprehensive guide provides complete coverage of every aspect of the building technologist's profession It details design and installation procedures describes all relevant equipment and hardware and illustrates the preparation of working drawings and construction details that meet project specifications code requirements and industry standards The author establishes procedures for professional field inspections and equipment operations tests provides real world examples from both residential and nonresidential construction projects and makes specific references to code compliance throughout the text This new edition incorporates changes in building codes advances in materials and design techniques and the emergence of computer aided design CAD while retaining the logical structure and helpful special features of the first edition More than 1 100 drawings tables and photographs complement and illustrate discussions in the text Topics covered include Heating ventilating and air conditioning systems equipment and design Plumbing systems equipment and design Electrical and lighting systems equipment and design Testing adjusting and balancing procedures for all building systems Every aspect of the building technologist's profession from the creation of working drawings through on site supervision and systems

maintenance Extensive appendices include conversion factors duct design data test report forms for use in field work design forms and schedules for electrical HVAC and plumbing work and more **Piping Calculations Manual** Shashi Menon,2004-12-10 This on the job resource is packed with all the formulas calculations and practical tips necessary to smoothly move gas or liquids through pipes assess the feasibility of improving existing pipeline performance or design new systems Contents Water Systems Piping Fire Protection Piping Systems Steam Systems Piping Building Services Piping Oil Systems Piping Gas Systems Piping Process Systems Piping Cryogenic Systems Piping Refrigeration Systems Piping Hazardous Piping Systems Slurry and Sludge Systems Piping Wastewater and Stormwater Piping Plumbing and Piping Systems Ash Handling Piping Systems Compressed Air Piping Systems Compressed Gases and Vacuum Piping Systems Fuel Gas Distribution Piping Systems *Domestic Engineering and the Journal of Mechanical Contracting* ,1927 Diseases of Foreign Forest Trees Growing in the United States Perley Spaulding,1958 The term trees is used in the broad sense and includes trees shrubs and woody vines **Crane Valve World** ,1913 The Metal Worker, Plumber, and Steam Fitter ,1911 **United Association of Journeymen and Apprentices of the Plumbing and Pipe Fitting Industry Journal** ,1964 Merchant Plumber and Fitter ,1915 Power ,1926 **American Society of Heating and Ventilating Engineers Guide** American Society of Heating and Air-Conditioning Engineers,1925 **Heating, Ventilating, Air Conditioning Guide** ,1925 **Domestic Engineering** ,1915 Science Abstracts ,1908 **The Valve World** ,1912 *The Valve World* Justin Wallace McEachren,1912 **Engineering Abstracts** Institution of Civil Engineers (Great Britain),1921 Engineering Abstracts from the Current Periodical Literature of Engineering and Applied Science, Published Outside the United Kingdom Institution of Civil Engineers (Great Britain),1921

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$ $\Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta H = 285 \text{ kJ/mole}$

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 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$

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web this activity provides a demonstration of hess law using three reactions the solubility NaOH in water the solubility NaOH in HCl and the reaction of a solution of HCl and a solution of NaOH

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web hess s law worksheet answers 1 calculate h for the reaction $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ from the following data $\text{C}_2\text{H}_4(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ $\Delta H = 1411 \text{ kJ}$ $\text{C}_2\text{H}_6(\text{g}) + 3\frac{1}{2}\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 3\text{H}_2\text{O}(\text{l})$ $\Delta H = 1560 \text{ kJ}$

7a first law enthalpy calorimetry and hess s law worksheet - Oct 05 2022

web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

hess law lab science notes - Apr 30 2022

web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kj in your final answer multiply the mass by the change in temperature and the c given trial 1 use hess s law and the following equations and ΔH values to determine the heat of reaction for the reaction C_2H_4

experiment 9 enthalpy of reaction hess s law - Jul 14 2023

web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

hess s law lab worksheet hess s law enthalpy of reaction studocu - Nov 06 2022

web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

hess s law lab report thermodynamics and calorimetry - Dec 07 2022

web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 2 $\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 3