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Rhetorical Machines John Jones, Lavinia Hirsu, 2019-07-02 A landmark volume that explores the interconnected nature of technologies and rhetorical practice *Rhetorical Machines* addresses new approaches to studying computational processes within the growing field of digital rhetoric While computational code is often seen as value neutral and mechanical this volume explores the underlying and often unexamined modes of persuasion this code engages In so doing it argues that computation is in fact rife with the values of those who create it and thus has powerful ethical and moral implications From Socrates's critique of writing in Plato's *Phaedrus* to emerging new media and internet culture the scholars assembled here provide insight into how computation and rhetoric work together to produce social and cultural effects This multidisciplinary volume features contributions from scholar practitioners across the fields of rhetoric computer science and writing studies It is divided into four main sections *Emergent Machines* examines how technologies and algorithms are framed and entangled in rhetorical processes *Operational Codes* explores how computational processes are used to achieve rhetorical ends *Ethical Decisions and Moral Protocols* considers the ethical implications involved in designing software and that software's impact on computational culture and the final section includes two scholars responses to the preceding chapters Three of the sections are prefaced by brief conversations with chatbots autonomous computational agents addressing some of the primary questions raised in each section At the heart of these essays is a call for emerging and established scholars in a vast array of fields to reach interdisciplinary understandings of human machine interactions This innovative work will be valuable to scholars and students in a variety of disciplines including but not limited to rhetoric computer science writing studies and the digital humanities

Who Can You Trust? Rachel Botsman, 2017-11-14 If you can't trust those in charge who can you trust From government to business banks to media trust in institutions is at an all time low But this isn't the age of distrust far from it In this revolutionary book world renowned trust expert Rachel Botsman reveals that we are at the tipping point of one of the biggest social transformations in human history with fundamental consequences for everyone A new world order is emerging we might have lost faith in institutions and leaders but millions of people rent their homes to total strangers exchange digital currencies or find themselves trusting a bot This is the age of distributed trust a paradigm shift driven by innovative technologies that are rewriting the rules of an all too human relationship If we are to benefit from this radical shift we must understand the mechanics of how trust is built managed lost and repaired in the digital age In the first book to explain this new world Botsman provides a detailed map of this uncharted landscape and explores what's next for humanity

Parochial and Family History of the Parish of Egloshayle in the County of Cornwall Sir John Maclean, 1871 *The American Historical Review* John Franklin Jameson, Henry Eldridge Bourne, Robert Livingston Schuyler, 1904 *American Historical Review* is the oldest scholarly journal of history in the United States and the largest in the world Published by the American Historical Association it covers all areas of historical research

Papers and Proceedings Hampshire Field Club

and Archaeological Society (Southampton, England),1914 **Survey of the Lands of William First Earl of Pembroke** Sir William Herbert,Roxburghe Club,1909 Digital Innovations for Mass Communications Paul Martin Lester,2013-11-12 In every field of mass communications advertising entertainment studies journalism public relations radio television film tourism and visual reporting professionals understand the importance of storytelling Regardless of whether the finished product is a commercial an in depth investigative piece a public service campaign an independent documentary a travelogue or a collection of photographs effective storytelling requires a combination of creativity empathy and expertise Through the innovative technologies and techniques described in this textbook students will learn how to turn passive readers and viewers into engaged and regular users The sixteen chapters each include a brief introduction assignments simple to follow step by step exercises and sources for additional information in which users will learn to produce apps informational graphics quick response codes quizzes simulations smartphone and table icons social media campaigns three dimensional pictures and video Students will work with the following programs Blogger Dreamweaver Excel Facebook GeoCommons Google Maps Illustrator Imgur iMovie Infogram iShowU JavaScript JustGive Kaywa Kickstarter LinkedIn Onvert Photoshop Pixel Resort QuickTime Reddit Second Life SurveyMonkey TheAppBuilder Twitter Vizualize Wikipedia Word WordPress and YouTube When digital innovations are added to traditional print and screen presentations a media user is not only allowed to interact with the information but can also physically engage with the story displayed Giving students the tools they need to transform their storytelling in this manner is the ultimate goal of this textbook Publications of the Catholic Record Society Catholic Record Society (Great Britain),1964 Vol 5 7 9 11 12 15 17 24 26 41 48 52 include Report of the Society 1907 1925 1927 1957 58 **Mastering Sync for Reddit Pro: The Ultimate Guide to Optimizing Your Reddit Experience** Navneet Singh, Table of Contents Introduction What is Sync for Reddit Pro Why Choose Sync for Reddit Pro Over Other Reddit Clients How This Book Will Help You Getting Started with Sync for Reddit Pro Installing Sync for Reddit Pro Setting Up Your Reddit Account Overview of the Interface Key Features of Sync for Reddit Pro Navigating Reddit with Sync for Reddit Pro Home Feed and Subreddits Subreddit Discovery and Exploration Sorting Posts Hot New Top and More Managing Reddit Notifications and Subscriptions Customization and Personalization Customizing the Theme Dark Mode Light Mode and Themes Organizing Subreddits into Groups Managing Multiple Accounts Syncing Across Devices Advanced Features The Power of Advanced Filters Using Mod Tools for Subreddit Moderators Post Management Saving Sharing and Voting Custom Actions and Gestures Maximizing Efficiency with Sync for Reddit Pro Keyboard Shortcuts for Power Users Multi Window Support and Multi Tasking Offline Mode and Data Usage Optimization Searching with Precision Subreddit Search Post Search etc Syncing and Privacy Features Syncing Data Across Devices with Cloud Storage Privacy Options and Settings Managing Reddit Permissions Troubleshooting and Support Common Issues and How to Solve Them Getting Help Sync for Reddit Pro s Support Channels FAQ Frequently Asked Questions Tips and Tricks for Pro Users Hidden Features and Easter

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Reports Great Britain. Royal Commission on Historical Manuscripts,1881 **Dēmosthenous pros Leptinēn. A revised text, with an intr. and notes by J.E. Sandys** Demosthenes,1890 The Classical Review ,1910 This companion to the Classical Quarterly contains reviews of new work dealing with the literatures and civilizations of ancient Greece and Rome Over 300 books are reviewed each year *Proceedings of the Somersetshire Archaeological and Natural History Society* Somersetshire Archaeological and Natural History Society,1876 *Proceedings* Somersetshire Archaeological and Natural History Society,1876

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web limiting reactant theoretical and percentage yields key $4 KO_2 + 2 H_2O \rightarrow 4 KOH + 3 O_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C₃H₈ what is the percent yield
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$
 18g H₂O / 1 mole H₂O = 149.4g H₂O is your theoretical yield so percent yield = $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting_reagents_theoretical_actual_and_percent_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH₃ can be produced theoretically from the reaction of 5.0 g of N
 what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO}_2 + 3\text{H}_2\text{O}$
 if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO₂ would be produced
 step one figure out how many moles of the limiting reagent you have
 $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design
 yield = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $\frac{160 \text{ g Al}}{27 \text{ g/mol}} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$
 theoretical 302.3 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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 $80\text{ g I}_2\text{O}_5 \times \frac{1\text{ mol I}_2\text{O}_5}{1\text{ mol I}_2\text{O}_5} \times \frac{1\text{ mol I}_2}{1\text{ mol I}_2\text{O}_5} \times \frac{333.8\text{ g I}_2}{1\text{ mol I}_2\text{O}_5} = 28\text{ g CO} \times \frac{1\text{ mol CO}}{1\text{ mol CO}}$

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