

Chapter 17: Chemical Equilibrium

Worksheet: K_c and K_p

1. For the reaction $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$, $K_c = 1.7 \times 10^3$ at 1000 K.

- 1. Calculate K_p for this reaction at 1000 K.
- 2. For the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$, $K_c = 0.50$ at 400°C. Calculate K_p at this temperature.
- 3. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 4. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 5. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 6. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 7. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 8. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 9. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.
- 10. For the reaction $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$, $K_c = 4.0 \times 10^{-4}$ at 1500 K. Calculate K_p at this temperature.

Answers

- 1. $K_p = 1.7 \times 10^3 \times (RT)^{-1} = 1.7 \times 10^3 \times (0.0821 \times 1000)^{-1} = 2.1 \times 10^2$
- 2. $K_p = K_c (RT)^{-2} = 0.50 \times (0.0821 \times 673)^{-2} = 1.3 \times 10^{-4}$
- 3. $K_p = K_c (RT)^{2} = 4.0 \times 10^{-4} \times (0.0821 \times 1500)^2 = 6.1 \times 10^{-3}$
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- 9. $K_p = K_c (RT)^{2} = 4.0 \times 10^{-4} \times (0.0821 \times 1500)^2 = 6.1 \times 10^{-3}$
- 10. $K_p = K_c (RT)^{2} = 4.0 \times 10^{-4} \times (0.0821 \times 1500)^2 = 6.1 \times 10^{-3}$

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**Alison Warr, Penny Slingerland, Cathy
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