

Chemistry: Measurement: Problem Sheet 2

Warning: There are 100 marks available for this problem sheet. Please give your answers according to the instructions below.

1. Mass analysis experiment according to the following chemical equation:



(A. How many moles of silver are needed to produce 0.01 moles of barium?)

$$x \text{ moles AgNO}_3 + y \text{ moles BaCl}_2 \rightarrow z \text{ moles AgCl} + w \text{ moles Ba(NO}_3)_2$$

(B. From the chemical equation you given in Part A, how many moles of silver chloride are produced?

$$x \text{ moles AgNO}_3 + y \text{ moles BaCl}_2 \rightarrow z \text{ moles AgCl} + w \text{ moles Ba(NO}_3)_2$$

(C. From the chemical equation you have given in Part B, how many moles of water will be produced?

$$x \text{ moles AgNO}_3 + y \text{ moles BaCl}_2 \rightarrow z \text{ moles AgCl} + w \text{ moles Ba(NO}_3)_2$$

(D. From the chemical equation you have given in Part B, how many moles of nitrogen molecules will be made?

$$x \text{ moles AgNO}_3 + y \text{ moles BaCl}_2 \rightarrow z \text{ moles AgCl} + w \text{ moles Ba(NO}_3)_2$$

2. Molecular formula question: $\text{C}_2\text{H}_4\text{Cl}_2 + \text{AgNO}_3 \rightarrow \text{C}_2\text{H}_4\text{Cl}_2 + \text{AgNO}_3$

(A. How many moles of hydrogen chloride are required to produce 0.01 moles of silver?)

$$x \text{ moles C}_2\text{H}_4\text{Cl}_2 + y \text{ moles AgNO}_3 \rightarrow z \text{ moles C}_2\text{H}_4\text{Cl}_2 + w \text{ moles AgNO}_3$$

(B. How many moles of hydrogen chloride are required to produce 0.01 moles of silver?

$$x \text{ moles C}_2\text{H}_4\text{Cl}_2 + y \text{ moles AgNO}_3 \rightarrow z \text{ moles C}_2\text{H}_4\text{Cl}_2 + w \text{ moles AgNO}_3$$

(C. How many moles of water are produced from 0.01 moles of silver and water?

$$x \text{ moles C}_2\text{H}_4\text{Cl}_2 + y \text{ moles AgNO}_3 \rightarrow z \text{ moles C}_2\text{H}_4\text{Cl}_2 + w \text{ moles AgNO}_3$$

3. Finding the mass of chlorine gas produced when 0.01 moles of dichloroethane is heated



$$x \text{ moles C}_2\text{H}_4\text{Cl}_2 \rightarrow y \text{ moles C}_2\text{H}_4 + z \text{ moles Cl}_2$$

Answer: (A) 0.01 moles (B) 0.005 moles (C) 0.0025 moles (D) 0.00125 moles

Chemistry Stoichiometry Problem Sheet 2 Answer Key

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