

*Matter and Change, pp. 1-4*

- 2.
- 3.
- 4.
9. mass and volume
10. group or family
11. homogeneous
12. products
13. metal
14. solid state
15. chemical properties
16. atom
17. compound
18. nonmetal
19. mixture
20. pure substance
21. basic
22. metalloid
23. reactants: carbon and oxygen; product: carbon dioxide
24. reactant: mercury(II) oxide; products: mercury and oxygen
25. f
26. g
27. c
28. a
29. b
30. i
31. d
32. h
33. e
34. physical change
35. chemical change
36. physical change
37. physical change
38. chemical change
39. physical change
40. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
41. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
42. Although energy can be absorbed or released in a

5. c
6. c
7. c
8. a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

*Measurements and Calculations, pp. 5-8*

- a
2. c
3. c
4. a
9. time
10. mass
11. density
12. energy
13. length or distance
14. volume
15. area
16. qualitative
17. quantitative
18. qualitative
19. quantitative
20.  $3.00 \times 10^7$  km/s
21. three
22. 0.026 g or  $2.6 \times 10^{-2}$  g
23.  $2.5 \times 10^{-4}$  L
24. quantity
25. derived units
26. conversion factor
27. 0.0432 kg or  $4.32 \times 10^{-2}$  kg
28. 5400 mL or  $5.4 \times 10^3$  mL
29. 300. K
30.  $1.05 \times 10^3$  J
31.  $3.51 \times 10^3$  cm
32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
35. The graph of two quantities that are directly proportional is a straight line through the origin.

5. c
6. c
7. a
8. a

# Chemistry Standardized Test Practice Answer Key Matter

**Karin Nielsen-Saines**



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My goal in writing this text was to relate the fundamental concepts of general organic and biological chemistry to the world around us and in this way illustrate how chemistry explains many aspects of everyday life A key feature is the use of molecular art to illustrate and explain common phenomena we encounter every day Each topic is broken down into small chunks of information that are more manageable and easily learned Students are given enough detail to understand basic concepts such as how soap cleans away dirt and why trans fats are undesirable in the diet without being overwhelmed

*Holt Chemistry* Ralph Thomas Myers,2004

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