

QUESTION 1

Volume of pipette used 25.00 cm^3
 (V₀)

Titration	Rough	1	2	3
Final burette reading (cm ³)	20.80	41.40	20.40	40.90
Initial burette reading (cm ³)	0.00	20.80	0.00	20.40
Volume of A used	20.80	20.60	20.40	20.50

② ^{average} Volume of A used = $\frac{20.60 + 20.40 + 20.50}{3}$
 $= 20.50 \text{ cm}^3$

③.1) Concentration of A in mol/dm³ = $\frac{\text{conc. in g/dm}^3}{\text{molar mass}}$
 $= \frac{4.4 \text{ g/dm}^3}{2[13] + 32 + 4[16]}$
 $= \frac{4.4}{98} = 0.0449 \text{ mol/dm}^3$

③.2) Concentration of B in mol/dm³, C_B

Using $\frac{C_A V_A}{C_B V_B} = \frac{n_A}{n_B}$

$$\frac{0.0449 \times 20.50}{C_B \times 25} = \frac{1}{1}$$

$$C_B = \frac{0.0449 \times 20.50}{25} = 0.03287 \text{ mol/dm}^3$$

Chemistry Practical Question And Answer

Kathleen Armour



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Curricula Designs Quality of Teaching and Learning Quality and Academic Assessment The School University of the Future Challenges for Higher Education in the 21st Century New Managerial Models for Education Financing the New Model for Education of the 21st Century The Quality Milestones for Education of the 21st Century Evaluation in Academia The Role of Teachers International Collaborations for Joint Programs Degrees Industry Academia Synergies Research Laboratories Management **The International Baccalaureate Diploma Programme** Tim Pound,2006-05-02 Introducing a balanced look at the experience of implementing and teaching the increasingly respected qualification the International Baccalaureate this book is a rich resource for all teachers school leaders and managers involved with or considering the qualification

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