

**2.1| INTRODUCTION**

In 1809 that John Dalton gave his atomic theory known as Dalton's atomic theory. Dalton considered atom to be indivisible, and indestructible.

Towards the end of 19th century, however, it came to known that atom consists of subatomic particles such as electrons, protons and neutrons.

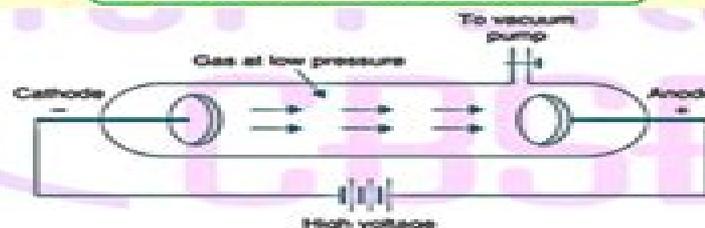
It mean atom was no more indivisible.

**SUBATOMIC PARTICLES**

- 1) Discovery of Electrons,
- 2) Protons and
- 3) Neutrons

**2.2| DISCOVERY OF ELECTRON (CATHODE RAYS): DISCHARGE TUBE EXP.**

- (i) A discharge tube is a cylindrical hard glass tube.
- (ii) Fitted with two metallic electrodes.
- (iii) The tube has a side-tube with a stopcock.
- (iv) The discharge tube can be evacuated or filled with any gas to any pressure.
- (v) A high potential difference of 10000V was applied across the electrodes.

**WORKING OF EXPERIMENT**

- (a) A discharge tube containing air under normal pressure. Electricity does not flow through the gas.
- (b) When the pressure inside tube is reduced to about  $10^{-2}$  atm, the gas starts conducting electricity, and it starts emitting light.

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**Karin Nielsen-Saines**



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