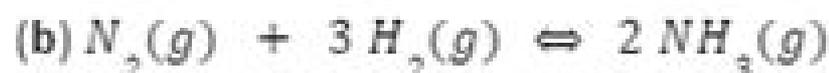


1. Write equilibrium constant expressions,  $K_p$ , for the following reactions.



$$\text{Ans. } K_p = \frac{P_{\text{CO}_2} \times P_{\text{H}_2}}{P_{\text{H}_2\text{O}} \times P_{\text{CO}}}$$



$$\text{Ans. } K_p = \frac{P_{\text{NH}_3}^2}{P_{\text{H}_2}^3 \times P_{\text{N}_2}}$$

2. A mixture containing equal concentrations of methane and steam is passed over a nickel catalyst at 1000 K. The emerging gas has the composition  $[\text{CO}] = 0.1027 \text{ mol dm}^{-3}$ ,  $[\text{H}_2] = 0.3080 \text{ mol dm}^{-3}$ , and  $[\text{CH}_4] = [\text{H}_2\text{O}] = 0.8973 \text{ mol dm}^{-3}$ . Assuming this mixture is at equilibrium, calculate the equilibrium constant  $K_c$  for the reaction

$$\begin{aligned} \text{Ans. } K_c &= \frac{[\text{CO}][\text{H}_2]^3}{[\text{CH}_4][\text{H}_2\text{O}]} = \frac{0.1027 \text{ mol dm}^{-3} \times (0.3080 \text{ mol dm}^{-3})^3}{0.8973 \text{ mol dm}^{-3} \times 0.8973 \text{ mol dm}^{-3}} \\ &= 3.727 \times 10^{-3} \text{ mol}^2 \text{ dm}^{-6} \end{aligned}$$

# Chemical Equilibrium Problems With Solutions

**John T. Donoghue**



## **Chemical Equilibrium Problems With Solutions:**

**A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter s main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings *Principles and Applications of Aquatic Chemistry* François M. M. Morel, Janet G. Hering, 1993-03-10 Presents aquatic chemistry in a way that is truly useful to those with diverse backgrounds in the sciences Major improvements to this edition include a complete rewrite of the first three background chapters making them user friendly There is less emphasis on mathematics and concepts are illustrated with actual examples to facilitate understanding **Exploring Chemical Analysis** Daniel C. Harris, 2005 Exploring Chemical Analysis teaches students how to understand analytical results and how to use quantitative manipulations preparing them for the problems they will encounter **Surface Complexation Modeling** David A. Dzombak, François M. M. Morel, 1991-01-16 Provides a description of the thermodynamic model data treatment procedures and the thermodynamic constants for hydrous ferric

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