

20 Points

Name: Phyllis A. Smith Lab Section: 200

Suppose that you had the correct response for the following composition for liquid state nitrogen boiling point unknown:

CH₂N₂O

C₂H₆

NO₂

H₂O

100g
Liquid

100.000g

100.00g

100g
Liquid

1. If you have these conditions at a 1.00 bar total pressure, what composition would be predicted by assumption of water to condense at 17.0°C? (20 points)

$$\frac{100g \text{ Liquid}}{100g} \left| \frac{100g}{100g} \right| \frac{17}{17.0} \left| \frac{100g}{100g} \right| = 17.0 \text{ mol } H_2O$$

17.0 mol

2. If you were in the mixture and wanted the same composition of NO₂ and H₂O, what would be the boiling point of the mixture at 1.00 bar? (20 points)

$$M_1 = \frac{10g \text{ sugar}}{100g \text{ water}} \left| \frac{100g}{100g} \right| \frac{1.00 \text{ bar}}{1.00 \text{ bar}} \left| \frac{100g}{100g} \right| = 10.0 \text{ mol } H_2O$$

$$M_2 = \frac{10g \text{ sugar}}{80g \text{ H}_2\text{O}} \left| \frac{100g}{100g} \right| \frac{1.00 \text{ bar}}{1.00 \text{ bar}} = 10.0 \text{ mol } H_2O$$

10.0 mol

10.0 mol

10.0 mol

10.0 mol

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Scott C. Dulebohn



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