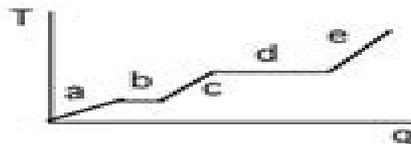


6. The highest vapor pressure is observed for which of the following liquid/temperature combinations?
- $C_3H_8$  at 270 K
  - $C_2H_6$  at 300 K
  - $C_3H_{12}$  at 300 K
  - $HOC_6H_5OH$  at 300 K
  - $HOC_6H_5OH$  at 270 K



7. Which regions on the on the heating curve shown (Temperature versus heat, "q") corresponds to
- a pure gas increasing in temperature **e**
  - a pure liquid increasing in temperature **c**
  - a solid increasing in temperature **a**
  - a solid melting **b**
  - a liquid boiling **d**
8. On the above heating curve, which phases are present:
- In region "a" **solid**
  - In region "b" **solid/liquid**
  - In region "c" **liquid**
  - In region "d" **liquid/gas**
  - In region "e" **gas**
9. A solid is soft, and has a low melting point (below 300°C). Its solid, molten and aqueous forms are all nonconductors. Classify the solid.
10. Indicate the type of crystalline solid each of the following would form upon solidification. Identify the type of particles found at the lattice points and the types of attractive forces.

	Type of crystalline solid	Type of particles at lattice points	IMF
(a)	<b>molecular solid</b>	<b>atoms</b>	<b>metallic bond</b>
(b)	<b>molecular</b>	<b>molecules</b>	<b>hydrogen bonds</b>
(c)	<b>molecular</b>	<b>molecules</b>	<b>LDF</b>
graphite	<b>covalent network</b>	<b>atoms</b>	<b>covalent bond</b>

# Chapter Section 12 1 Liquids And Solids Answer Key

**Ralph Thomas Myers**



## Chapter Section 12 1 Liquids And Solids Answer Key:

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Instructor's Manual to Accompany Introductory College Physics Joseph Francis Mulligan,1985      **Te HS&T 2007 Shrt Crs M** Holt Rinehart & Winston,2007      *Holt Science: Teacher's edition* Joseph Abruscato,1989      *Reading in the Content Area with Literature Connections* ,2000      *CHEMISTRY* SILBERBERG,2003      **Chemistry** John S. Phillips,Cheryl Wistrom,2000      *Holt Chemistry* Ralph Thomas Myers,2004      **Elementary Principles of Chemical Processes** Richard M. Felder,Ronald W. Rousseau,Lisa G. Bullard,2015-12-17 This best selling text prepares students to formulate and solve material and energy balances in chemical process systems and lays the foundation for subsequent courses in chemical engineering The text provides a realistic informative and positive introduction to the practice of chemical engineering

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