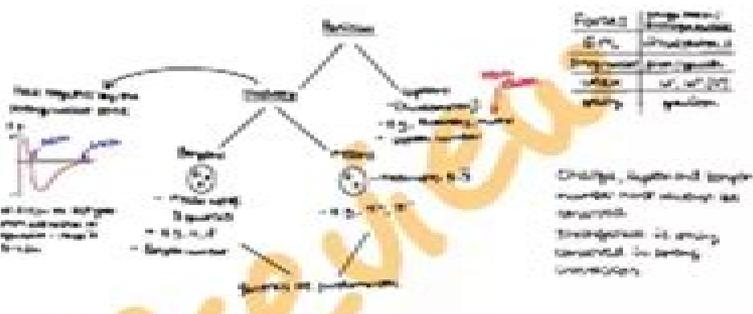


Particles and Quantum



Excitation: when a particle and its corresponding antiparticle collide and their rest energy is converted into photons.

Pair production: when a photon turns into a particle and its corresponding antiparticle. The photon must have at least the sum amount of energy to create the particles.



Excitation: Existing more than just excited enough energy to count their atoms.

Alpha particles: Helium nucleus, ^4_2He . Highly ionizing, so it is dangerous. Stopped by a piece of paper or foil on skin.

Beta: fast moving electron. $\rightarrow e^-$. Medium penetrating and velocity slowing. Stopped by a few cm of aluminium.

Gamma: high energy. EM ray or photon emitted from a high energy nucleus. Highly penetrating. Stopped by a few m of concrete.

Isotopes: Same element - same number of protons. Different relative atomic mass. Different number of neutrons.

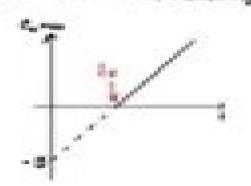
Specific charge: $\frac{Q}{m}$ (C/kg). Charge to mass ratio.

Electron rest mass: charge of an electron when stationary. 9.11×10^{-31} kg.



Photoelectric effect: use photons of sufficient energy on electrons on the surface of the metal, liberating them.

- $E_{\text{photon}} = hf = \phi + E_{\text{max}}$
- ϕ - photon absorbed - $h\nu$ (work function)
- E_{max} - frequency
- ϕ - work function - minimum energy to liberate electrons
- E_{photon} - energy that will give an e^- its maximum
- E_{max} - stopped frequency - minimum frequency needed for electrons to be liberated



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons is 0 if the frequency is less than the threshold frequency. The intensity of light cannot liberate the e^- if the frequency is below the threshold but it will liberate the number of electrons that is proportional to the power.

Stopping potential: the pd needed to stop any electrons from reaching the other side.

- $E_{\text{photon}} = hf = \phi + eV_s$
- eV_s - charge of electron
- V_s - stopping potential

De Broglie: The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction of a crystal. This is because the electron is diffracting which is wave-like.



$\lambda = \frac{h}{p} = \frac{h}{mv}$

λ - de Broglie wavelength
 p - momentum / mass x velocity

Fluorescent tubes:

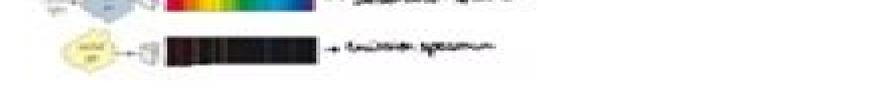


The tube contains mercury under a high voltage is applied. This excites the gas resulting in which some of the mercury atoms emit photons when they relax. When the gas relaxes it emits UV light which is higher energy than what the electrons have, so the photons in the UV range. A phosphor coating in the tube absorbs the photons and emit visible light.

Energy levels:

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In a gas discharge tube or in a gas discharge cell when it is giving the electron energy.

Ionization energy: is enough energy to give an electron, so will completely remove the electron.



A Level Physics Notes

H Kauffman



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