

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{56}_{26}\text{Fe}$	B2																
3 (b)	$^{23}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> <th>charge</th> </tr> </thead> <tbody> <tr> <td><math>^{17}\text{O}</math></td> <td>8</td> <td>9</td> <td>10</td> <td>2-</td> </tr> <tr> <td><math>^{27}\text{Al}</math></td> <td>13</td> <td>14</td> <td>10</td> <td>3+</td> </tr> </tbody> </table>		protons	neutrons	electrons	charge	$^{17}\text{O}$	8	9	10	2-	$^{27}\text{Al}$	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
$^{17}\text{O}$	8	9	10	2-														
$^{27}\text{Al}$	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table border="1"> <thead> <tr> <th></th> <th>protons</th> <th>neutrons</th> <th>electrons</th> </tr> </thead> <tbody> <tr> <td><math>^{32}\text{S}</math></td> <td>16</td> <td>16</td> <td>16</td> </tr> <tr> <td><math>^{32}\text{S}^{2-}</math></td> <td>16</td> <td>16</td> <td>18</td> </tr> </tbody> </table>		protons	neutrons	electrons	$^{32}\text{S}$	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
$^{32}\text{S}$	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2\text{(SO}_4\text{)}_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

# A Level Chemistry And Answers Key

**Michael Rajnik**



## **A Level Chemistry And Answers Key:**

## Reviewing **A Level Chemistry And Answers Key**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**A Level Chemistry And Answers Key**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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### **Table of Contents A Level Chemistry And Answers Key**

1. Understanding the eBook A Level Chemistry And Answers Key
  - The Rise of Digital Reading A Level Chemistry And Answers Key
  - Advantages of eBooks Over Traditional Books
2. Identifying A Level Chemistry And Answers Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an A Level Chemistry And Answers Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from A Level Chemistry And Answers Key
  - Personalized Recommendations
  - A Level Chemistry And Answers Key User Reviews and Ratings
  - A Level Chemistry And Answers Key and Bestseller Lists

5. Accessing A Level Chemistry And Answers Key Free and Paid eBooks
  - A Level Chemistry And Answers Key Public Domain eBooks
  - A Level Chemistry And Answers Key eBook Subscription Services
  - A Level Chemistry And Answers Key Budget-Friendly Options
6. Navigating A Level Chemistry And Answers Key eBook Formats
  - ePub, PDF, MOBI, and More
  - A Level Chemistry And Answers Key Compatibility with Devices
  - A Level Chemistry And Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of A Level Chemistry And Answers Key
  - Highlighting and Note-Taking A Level Chemistry And Answers Key
  - Interactive Elements A Level Chemistry And Answers Key
8. Staying Engaged with A Level Chemistry And Answers Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers A Level Chemistry And Answers Key
9. Balancing eBooks and Physical Books A Level Chemistry And Answers Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection A Level Chemistry And Answers Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine A Level Chemistry And Answers Key
  - Setting Reading Goals A Level Chemistry And Answers Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of A Level Chemistry And Answers Key
  - Fact-Checking eBook Content of A Level Chemistry And Answers Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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