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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o₂ o₂ 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen
2 moles leftover hydrogen 1 mole leftover

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web if you start with 14.82 g of Ca(OH)_2 and 16.35 g of H_2SO_4 a determine the limiting reagent b determine the number of moles of H_2O produced c determine the number of grams of CaSO_4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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