

Question 1: Kinetics of a Reaction and Equilibrium

1.1. Determining Reaction Order and Rate Law

- a. Second
- b. First
- c. Third
- d. Fourth
- e. Zeroth

1.2. Interpret the rate law for the reaction $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ and determine the overall order of the reaction.

1.3. For the reaction $\text{A} \rightarrow \text{B} + \text{C}$, the half-life of A is 10 minutes. How long will it take for A to decrease to 1/8 of its initial concentration?

- a. 10 min
- b. 20 min
- c. 30 min
- d. 40 min
- e. 50 min
- f. 60 min
- g. 70 min
- h. 80 min

1.2. Determining Reaction Order and Equilibrium

- a. $\text{rate} = k[\text{A}]^2[\text{B}]$
- b. $\text{rate} = k[\text{A}][\text{B}]^2$
- c. $\text{rate} = k[\text{A}][\text{B}]$

1.3. For the reaction $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$, the equilibrium constant K_c is 4.0. If the initial concentrations of A and B are 1.0 M and 2.0 M, respectively, what is the equilibrium concentration of C ?

- a. 0.5 M
- b. 1.0 M
- c. 1.5 M
- d. 2.0 M
- e. 2.5 M
- f. 3.0 M
- g. 3.5 M
- h. 4.0 M

1.4. For the reaction $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$, the equilibrium constant K_c is 4.0. If the initial concentrations of A and B are 1.0 M and 2.0 M, respectively, what is the equilibrium concentration of D ?

1.3. Calculating Reaction Order and Equilibrium

- a. Second order

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Marco Casella



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