

# Chemistry

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(Chapter 13) (Amines) (Exercises)  
(Class 12)

## Question 13.1:

Write IUPAC names of the following compounds and classify them into primary, secondary and tertiary amines.



## Answer 13.1:

(i) 1-Methylethanamine (1° amine)

(ii) Propan-1-amine (1° amine)

(iii) N-Methyl-2-methylethanamine (2° amine)

(iv) 2-Methylpropan-2-amine (1° amine)

(v) N-Methylbenzamine or N-methylaniline (2° amine)

(vi) N-Ethyl-N-methylethanamine (3° amine)

(vii) 3-Bromobenzenamine or 3-bromoaniline (1° amine)

## Question 13.2:

Give one chemical test to distinguish between the following pairs of compounds.

(i) Methylamine and dimethylamine

(ii) Secondary and tertiary amines

(iii) Ethylamine and aniline

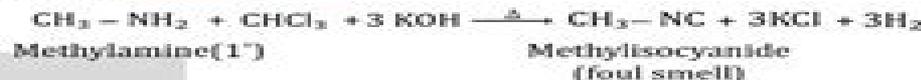
(iv) Aniline and benzylamine

(v) Aniline and N-methylaniline.

## Answer 13.2:

(i) Methylamine and dimethylamine can be distinguished by the carbylamine test.

**Carbylamine test:** Aliphatic and aromatic primary amines on heating with chloroform and ethanolic potassium hydroxide form foul-smelling isocyanides or carbylamines. Methylamine (being an aliphatic primary amine) gives a positive carbylamine test, but dimethylamine does not.



(ii) Secondary and tertiary amines can be distinguished by allowing them to react with Hinsberg's reagent (benzenesulphonyl chloride, C<sub>6</sub>H<sub>5</sub>SO<sub>2</sub>Cl). Secondary amines react with Hinsberg's reagent to form a product that is insoluble in an alkali. For example, N, N-diethylamine reacts with Hinsberg's reagent to form N, N-diethylbenzenesulphonamide, which is insoluble in an alkali. Tertiary amines, however, do not react with Hinsberg's reagent.



(iii) Ethylamine and aniline can be distinguished using the azo-dye test. A dye is obtained when aromatic amines react with HNO<sub>2</sub> (NaNO<sub>2</sub> + dil.HCl) at 0-5°C, followed by a reaction with the alkaline solution of 2-naphthol. The dye is usually yellow, red, or orange in colour. Aliphatic amines give a brisk effervescence due (to the evolution of N<sub>2</sub> gas) under similar conditions.

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# Chemistry Chapter 13 Solutions

**J. Philip Bromberg**



## Chemistry Chapter 13 Solutions:

**Solutions Manual to Accompany Organic Chemistry** Jonathan Clayden, Stuart Warren, Stuart G. Warren, 2013 This text contains detailed worked solutions to all the end of chapter exercises in the textbook Organic Chemistry Notes in tinted boxes in the page margins highlight important principles and comments

**Introduction to Chemistry** Amos Turk, 2013-07-15 Introduction to Chemistry is a 26 chapter introductory textbook in general chemistry This book deals first with the atoms and the arithmetic and energetics of their combination into molecules The subsequent chapters consider the nature of the interactions among atoms or the so called chemical bonding This topic is followed by discussions on the nature of intermolecular forces and the states of matter This text further explores the statistics and dynamics of chemistry including the study of equilibrium and kinetics Other chapters cover the aspects of ionic equilibrium acids and bases and galvanic cells The concluding chapters focus on a descriptive study of chemistry such as the representative and transition elements organic and nuclear chemistry metals polymers and biochemistry Teachers and undergraduate chemistry students will find this book of great value

**NCERT Class 12 Chemistry Solution** Jagran Josh, 2017-06-14 Students who will appear for the upcoming Class 12th Board exam pay attention Here we are with the most popular study material that will add highest grades in your marks sheet We have come up with chapter wise solution package for Chemistry that will help you to prepare in a smart way Students who are looking to save time and prepare effortlessly for the upcoming class 12th Board exam they must buy this comprehensive solution package of Chemistry subject This eBook comprises chapter wise solution to every question and also explains the concept behind it with in depth analysis It is a must have eBook for every student who wants to get good grades in the board exam and get admission in top colleges for further studies Key Features Chapter wise solution to every question of all 16 chapters Every question is solved in a step by step way for your better learning Best way to prepare and save a lot of time Effortless way to revise and get good grades in the exam

**Solutions Guide for Chemistry** Steven S. Zumdahl, Kenneth C. Brooks, 1989 *Fundamentals of Chemistry: A Modern Introduction (1966)* Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro s Law and Graham s Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and

basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry

Solutions Guide for Chemistry, Third Edition, by Steven S. Zumdahl Kenneth C. Brooks,1993 *Study Guide to Accompany Basics for Chemistry* Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

**Princeton Review MCAT General Chemistry Review, 5th Edition** The Princeton Review,2025-11-04 Everything students need to know to succeed on the General Chemistry section of the MCAT exam Medical College Admission Test including 3 full length practice tests With in depth coverage of all MCAT general chemistry topics The Princeton Review MCAT General Chemistry Review helps you practice your way to excellence featuring everything you need for a high score Thorough content review for of all MCAT general chemistry topics 3 full length online practice tests with detailed answer explanations and score reports Test taking strategies geared toward general chemistry mastery Bulleted chapter summaries for quick review Full color illustrations diagrams and tables End of chapter drills and explanations plus MCAT style passages and questions 3 scholarly journal articles with exercises to strengthen your critical analysis Extensive reference glossary The Princeton Review MCAT General Chemistry Review covers Chemistry Fundamentals Atomic Structure and Periodic Trends Bonding and Intermolecular Forces Thermodynamics Phases Gases Kinetics Equilibrium Acids and Bases Eletrochemistry MCAT Math for General Chemistry

*Student Solutions Manual for Chang's Chemistry* Philip C. Keller,Raymond Chang,Jill L. Keller,1988

**Physical Chemistry** Kurt W. Kolasinski,2016-09-07 Much of chemistry is

motivated by asking How do I make a primary alcohol React a Grignard reagent with formaldehyde Physical chemistry is motivated by asking Why The Grignard reagent and formaldehyde follow a molecular dance known as a reaction mechanism in which stronger bonds are made at the expense of weaker bonds If you are interested in asking why and not just how then you need to understand physical chemistry Physical Chemistry How Chemistry Works takes a fresh approach to teaching in physical chemistry This modern textbook is designed to excite and engage undergraduate chemistry students and prepare them for how they will employ physical chemistry in real life The student friendly approach and practical contemporary examples facilitate an understanding of the physical chemical aspects of any system allowing students of inorganic chemistry organic chemistry analytical chemistry and biochemistry to be fluent in the essentials of physical chemistry in order to understand synthesis intermolecular interactions and materials properties For students who are deeply interested in the subject of physical chemistry the textbook facilitates further study by connecting them to the frontiers of research Provides students with the physical and mathematical machinery to understand the physical chemical aspects of any system Integrates regular examples drawn from the literature from contemporary issues and research to engage students with relevant and illustrative details Important topics are introduced and returned to in later chapters key concepts are reinforced and discussed in more depth as students acquire more tools Chapters begin with a preview of important concepts and conclude with a summary of important equations Each chapter includes worked examples and exercises discussion questions simple equation manipulation questions and problem solving exercises Accompanied by supplementary online material worked examples for students and a solutions manual for instructors Fifteen supporting videos from the author presenting such topics as Entropy Rate Laws Sequestration Electrochemistry etc Written by an experienced instructor researcher and author in physical chemistry with a voice and perspective that is pedagogical and engaging

**Journal of the Chemical Society** Chemical Society (Great Britain), 1897 Titles of chemical papers in British and foreign journals included in Quarterly journal v 1 12

College Chemistry Steven Boone, Drew H. Wolfe, 2011-11-01 The Collins College Outline for College Chemistry is a comprehensive guide to the fundamental concepts behind chemical reactions bonding equilibria and thermodynamics with topics ranging from simple chemical measurements and the basics of atoms and molecules to entropy electrochemistry and nuclear chemistry Fully revised and updated by Dr Steven Boone College Chemistry includes practical test yourself sections with answers and complete explanations at the end of each chapter Also included are essential vocabulary definitions and sample exercises as well as detailed images charts and diagrams The Collins College Outlines are a completely revised in depth series of study guides for all areas of study including the Humanities Social Sciences Mathematics Science Language History and Business Featuring the most up to date information each book is written by a seasoned professor in the field and focuses on a simplified and general overview of the subject for college students and where appropriate Advanced Placement students Each Collins College Outline is fully integrated with

the major curriculum for its subject and is a perfect supplement for any standard textbook

**Solutions Manual to Accompany Physical Chemistry** J. Philip Bromberg, 1984

**5 Steps to a 5 AP Chemistry, 2012-2013 Edition** Richard H. Langley, John Moore, 2011-06-10 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That's why we've created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You'll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the Knowledge Step 5 Build Your Confidence Topics include Basics Reactions and Periodicity Stoichiometry Gases Thermodynamics Spectroscopy Light and Electrons Bonding Solids Liquids and Intermolecular Forces Solutions and Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry Organic Chemistry Experimental

Modern Scientific Knowledge of Nature, Man, and Society Frederick Albert Cleveland, 1929

*Chemistry Principles Student Solutions Guide* Steven S. Zumdahl, 2001-08 End of chapter problems with blue numbers appear in this guide These answers also appear at the end of the text These solutions include a mix of odd and even numbered problems comprising about half of the total problem sets found in the text

**Chemistry, Student Solutions Manual** James N. Spencer, George M. Bodner, Lyman H. Rickard, 2008-01-09 This innovative text provides a 15 chapter introduction to the fundamental concepts of chemistry The material is then supplemented by special topics at the end of each chapter

The Chemical News, 1927

**CliffsNotes Chemistry Practice Pack** Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level

*Journal - Chemical Society, London* Chemical Society (Great Britain), 1897

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