

CHAPTER 9 REVIEW

Stoichiometry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- _____ The coefficients in a chemical equation represent the
 - masses in grams of all reactants and products.
 - relative number of moles of reactants and products.
 - number of atoms of each element in each compound in a reaction.
 - number of valence electrons involved in a reaction.
- _____ Which of the following would not be studied within the topic of stoichiometry?
 - the mole ratio of Al to Cl in the compound aluminum chloride
 - the mass of carbon produced when a known mass of sucrose decomposes
 - the number of moles of hydrogen that will react with a known quantity of oxygen
 - the amount of energy required to break the ionic bonds in CaF_2
- _____ A balanced chemical equation allows you to determine the
 - mole ratio of any two substances in the reaction.
 - energy released in the reaction.
 - electron configuration of all elements in the reaction.
 - reaction mechanism involved in the reaction.
- _____ The relative number of moles of hydrogen to moles of oxygen that react to form water represents a(n)
 - reaction sequence.
 - bond energy.
 - mole ratio.
 - element proportion.
- Given the reaction represented by the following unbalanced equation:
 $\text{N}_2\text{O}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 - Balance the equation.

b. What is the mole ratio of NO_2 to O_2 ?

c. If 20.0 mol of NO_2 form, how many moles of O_2 must have been consumed?

d. Twice as many moles of NO_2 form as moles of N_2O are consumed. True or False?

e. Twice as many grams of NO_2 form as grams of N_2O are consumed. True or False?

Chapter 9 Stoichiometry Answers

Robert Boikess



Chapter 9 Stoichiometry Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Chemistry Workbook For Dummies Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you ll build your chemistry skills and improve your performance both in and out of the science lab You ll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence *High School Chemistry* Michael A. Buben, 2018-05-18 The Teacher s Edition is designed to add direction to High School Chemistry which is an outline of notes The books go hand in hand Modified Second Edition as a student s workbook and guidance from the Teacher s Edition **Scientific Soapmaking** Kevin M. Dunn, 2010 Scientific Soapmaking bridges the gap between the technical and craft literature It explains the chemistry of fats oils and soaps and teaches sophisticated analytical techniques that can be carried out using equipment and materials familiar to makers of handcrafted soap *Answer Book to Accompany Boikess/Edelson, Chemical Principles, Third Edition* Robert S. Boikess, 1985

Student Solutions Manual, Chemistry, Principles & Practice, Second Edition John M. DeKorte, 1997

Instructor's Manual and Test Bank to Accompany Basic Concepts of Chemistry Malone, 2000-04-18

Programmed Problem-solving for First-year Chemistry Arnold B. Loebel, 1983 **Instructor's manual and solutions to problems to accompany General chemistry: principles and structures** James E. Brady, 1975-04

Solutions to Red Exercises - Chemistry Roxy Wilson, 2005-05 Prepared by Roxy Wilson of University of Illinois Urbana Champaign Full so lutions to all of the red numbered exercises in the text are pro vided Short answers to red exercises are found in the appendix of the text Solutions to Black Exercises Roxy Wilson, 1999-08 Solutions Manual to Accompany Chemistry Raymond Chang, 1984 **Solutions Manual to Accompany Chemical Principles** Robert Boikess, 1994

Student Solutions Manual to Accompany Chemistry & Chemical Reactivity, Fourth Edition, Kotz & Treichel

Alton J. Banks, Paul Treichel, 1999 **Solutions Manual to Accompany Boikess/Edelson, Chemical Principles, Third Edition** Robert S. Boikess, 1985 **Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers

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Solutions to Exercises in Chemistry, the Central Science, 2nd Edition Theodore L. Brown, Harold Eugene LeMay, 1981

Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Introduction to General Chemistry Jerome K. Holmes, 1969

Chemical Problem-solving by Dimensional Analysis Arnold B. Loebel, 1978

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