

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron(II) sulfide, the roman numeral tells us (2)
 - How many atoms of Fe are in one formula unit.
 - How many sulfide ions can be obtained by the ion atom.
 - The charge on each Fe ion.
 - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is (1)
 - Change the number of ions represented in the formula.
 - Change the charges on the other ions in the compound.
 - Change the formula so that it no longer represents the compound.
 - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{CH}_3\text{CO}_2\text{H}$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead(II) carbonate	PbCO_3
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Sulfuric acid	H_2SO_4
Nitrogen pentoxide	N_2O_5
Copper(II) carbonate	CuCO_3
Sodium sulfate	Na_2SO_4
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	HgSO_4
Nitrous acid	HNO_2
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron(II) chloride	FeCl_2
Iron(III) chloride	FeCl_3
Carbon(II) oxide	CO
Silver chloride	AgCl
Mercury perchlorate	Hg_2
Tetraphosphorus trisulfide	P_4S_3
Carbonic dioxide	CO_2

- Name the following compound and identify its molecular ions or acid.
 - $\text{Ca}(\text{OH})_2$, calcium hydroxide - II
 - $\text{Fe}(\text{NO}_3)_3$, Iron(III) nitrate - III
 - $\text{Fe}(\text{OH})_2$, Iron(II) hydroxide - II
 - N_2O_5 , Nitrogen pentoxide - III

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J Elliott



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