



COVALENT BONDING

Practice Problems

Do your answers to the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
 - NaCl
 - H_2O
 - HCl
 - KNO_3
 - H_2SO_4
 - H_2O_2
 - C_2H_6
 - CaCl_2
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
- Write the formula for the following molecular compounds:
 - nitrogen dioxide
 - carbon dioxide
 - carbon monoxide
 - nitrogen monoxide
 - nitrogen trioxide
 - nitrogen pentoxide
 - phosphorus pentoxide
 - phosphorus trioxide
 - phosphorus pentachloride
 - phosphorus trichloride
 - phosphorus pentasulfide
 - phosphorus trisulfide
 - phosphorus pentabromide
 - phosphorus tribromide
 - phosphorus pentaiodide
 - phosphorus triiodide
 - phosphorus pentaselenide
 - phosphorus triselenide
 - phosphorus pentatelluride
 - phosphorus ditelluride
 - phosphorus hexafluoride
 - phosphorus tetrafluoride
 - phosphorus hexachloride
 - phosphorus dichloride
 - phosphorus hexabromide
 - phosphorus dibromide
 - phosphorus hexaiodide
 - phosphorus diiodide
 - phosphorus hexasulfide
 - phosphorus disulfide
 - phosphorus hexatelluride
 - phosphorus ditelluride
 - phosphorus hexafluoride
 - phosphorus tetrafluoride
 - phosphorus hexachloride
 - phosphorus dichloride
 - phosphorus hexabromide
 - phosphorus dibromide
 - phosphorus hexaiodide
 - phosphorus diiodide
 - phosphorus hexasulfide
 - phosphorus disulfide
 - phosphorus hexatelluride
 - phosphorus ditelluride
- Write the name for each of the following molecular compounds:
 - CO_2
 - CO
 - NO_2
 - NO
 - N_2O_5
 - N_2O_3
 - P_2O_5
 - P_2O_3
 - P_2S_5
 - P_2S_3
 - P_2I_5
 - P_2I_3
 - P_2S_6
 - P_2S_2
 - P_2Te_6
 - P_2Te_2
 - P_2F_6
 - P_2F_4
 - P_2Cl_6
 - P_2Cl_2
 - P_2Br_6
 - P_2Br_2
 - P_2I_6
 - P_2I_2
 - P_2S_6
 - P_2S_2
 - P_2Te_6
 - P_2Te_2
 - P_2F_6
 - P_2F_4
 - P_2Cl_6
 - P_2Cl_2
 - P_2Br_6
 - P_2Br_2
 - P_2I_6
 - P_2I_2

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride. $\text{H}-\text{F}$
- Show covalent bonding in methane. C with four H atoms bonded to it.
- Show covalent bonding in carbon dioxide. $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene. $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules:
 - H_2O
 - H_2O_2
 - C_2H_6
 - C_2H_4
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_4\text{O}$
 - $\text{C}_2\text{H}_2\text{O}$
 - $\text{C}_2\text{H}_6\text{N}_2$
 - $\text{C}_2\text{H}_4\text{N}_2$
 - $\text{C}_2\text{H}_2\text{N}_2$
 - $\text{C}_2\text{H}_6\text{S}_2$
 - $\text{C}_2\text{H}_4\text{S}_2$
 - $\text{C}_2\text{H}_2\text{S}_2$
 - $\text{C}_2\text{H}_6\text{Se}_2$
 - $\text{C}_2\text{H}_4\text{Se}_2$
 - $\text{C}_2\text{H}_2\text{Se}_2$
 - $\text{C}_2\text{H}_6\text{Te}_2$
 - $\text{C}_2\text{H}_4\text{Te}_2$
 - $\text{C}_2\text{H}_2\text{Te}_2$
 - $\text{C}_2\text{H}_6\text{Br}_2$
 - $\text{C}_2\text{H}_4\text{Br}_2$
 - $\text{C}_2\text{H}_2\text{Br}_2$
 - $\text{C}_2\text{H}_6\text{I}_2$
 - $\text{C}_2\text{H}_4\text{I}_2$
 - $\text{C}_2\text{H}_2\text{I}_2$
 - $\text{C}_2\text{H}_6\text{Cl}_2$
 - $\text{C}_2\text{H}_4\text{Cl}_2$
 - $\text{C}_2\text{H}_2\text{Cl}_2$
 - $\text{C}_2\text{H}_6\text{F}_2$
 - $\text{C}_2\text{H}_4\text{F}_2$
 - $\text{C}_2\text{H}_2\text{F}_2$

SECTION 8.3 BONDING IN THE LITHIUM

- Write the Lewis structure for lithium hydride. LiH
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Brendan G. Carr



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